

Iodometric determination of glucose and ribose using potassium iodate as novel reagent in acetic acid medium

Abstract

In this work, we have reported a simple, cost effective and reliable method for the determination of glucose iodometrically by making use of potassium iodate. This volumetric method determines glucose instantly, thereby greatly reduces the time of determination.

Keywords: HIO_3 , glucose, iodometry, molecular weight, gluconic acid, distilled water, oxidation, hydroxyl group, sodium thiosulphate, metabolism, colorimetric, aromatic amines

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Introduction

Glucose is the major carbohydrate found in blood and a chief source of energy in human body. The blood glucose levels are perfectly maintained under the influence of hormones like insulin. However the hormonal imbalance sometimes may result in abnormalities of glucose metabolism and results in diseased condition.¹ Thus the detection of blood glucose levels can provide a basic understanding of the malfunctioning of the tissues and body. Although various methods are reported in the literature for the determination of glucose, few are commonly used via, colorimetric method, Benedict's method,² Bertrand's method,³ Hagedorn-Jenson method,⁴ chloramine-T⁵ method etc.

Literature survey revealed that combination of iodic acid and potassium iodide has been used for tri-methylsilylation of alcohols and phenols in the presence of HMDS and iodination of aromatic amines.⁶ Recently Kelsey B LaMartina et al.,⁷ showed that the combinations of ammonium iodate and catalytic *N*-hydroxyphthalimide (NHPI) in presence of acetic acid for the selective oxidation of *n*-butyl benzene directly to 1-phenylbutyl acetate in high yield⁷ and it is also reported as catalyst for the α -sulfonylation of enamines.⁸ Recently Rai et al.,⁹ used KIO_3 as a novel oxidizing agent for the synthesis of isoxazoles.⁹

In continuation of our work on synthetic and analytical applications of HIO_3 , we thought of an operationally simple titrimetric method for the determination of glucose. The method reported here makes use of the fact that glucose is known to undergo oxidation by HIO_3 , yielding gluconic acid involving one molecule of HIO_3 per molecule of glucose. From the mechanism shown below, it is evident that the reactive site involved for the attack of HIO_3 is the anomeric hydroxyl group. This moiety is more reactive than the other functional groups. Therefore in the presence of other materials HIO_3 first reacts with glucose and the rest needs more time to react under the experimental condition.

Materials and method

All reagents and chemicals used were of analytical reagent grade and were procured from SRL, India. Distilled water was used

throughout the experiment. In a typical experiment, a known excess of standard solution of HIO_3 was added to a known amount of glucose. After completion of the reaction, unreacted HIO_3 was determined by iodometry. By carrying out a blank experiment simultaneously, the amount of HIO_3 consumed was determined. As the overall reaction requires one mole of HIO_3 per molecule of glucose, which is equivalent to one mole of iodine and hence the molecular weight 'm' of glucose is determined using the following equation.

1 Mole of sugar \equiv 1 mol of $\text{I}_2 \equiv$ 2000 ml of 1N Sodium thiosulphate solution

$$M = \frac{2000 \times w}{(V_1 - V_2) \times N}$$

Where, m = Molecular weight of Glucose

W = Weight of the given sample

V_2 = Volume of sodium thiosulphate consumed (Blank)

V_1 = Volume of sodium thiosulphate consumed (experimental)

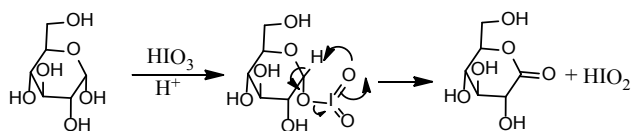
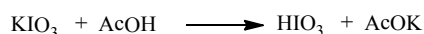
M = molarity of sodium thiosulphate

Determination of molecular weight of glucose/ribose: An accurately weighed (20-60mg) sample of glucose/ribose was dissolved in distilled water (10ml) in an Erlenmeyer flask. To this, a solution of 0.01mol of HIO_3 was introduced and it was heated to about 65°C on water bath, to this solution about 5ml of dilute sulphuric acid and 5ml of 10% potassium iodide was added and the liberated iodine was titrated against the standardized sodium thiosulphate solution using starch as indicator. In a similar way, a blank titration was conducted without adding glucose. From the difference in the volume of sodium thiosulphate solution consumed, the molecular weight 'm' was calculated using the above equation.

Results and discussions

The method reported here makes use of the fact that glucose is known to undergo an oxidative dehydrogenation by HIO_3 , yielding the gluconic acid by consuming one mole of HIO_3 per one molecule of

glucose or ribose (Scheme 1). Generally a known volume of HIO_3 is added to known mass of Glucose, after the completion of the reaction, the unreacted HIO_3 is determined iodometrically. By carrying out a parallel blank experiment the amount of the HIO_3 consumed is determined. As the overall reactions require one mole CAT per one mole of the Glucose, which is equivalent to mole of iodine, weight of the Glucose is determined by using equation shown above.



Scheme

Scheme I

Conclusion

We have developed a reliable, cost effective method for the determination of glucose using mild conditions and without the use of any sophisticated instruments and also this method requires short time.

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Conflicts of interest

The author declares that there are no conflicts of interest.

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