

# A review on the factors affecting the photocatalytic degradation of hazardous materials

## Abstract

The oxidation rates and efficiency of the photocatalytic system are highly dependent on a number of operational parameters that govern the photodegradation of the organic molecule. Several study have been reported the significance of operational parameter. The photodegradation depends on the some basic parameters which are concentration of substrate, amount of photocatalyst, pH of the solution, temperature of reaction medium, time of irradiation of light, the intensity of light, surface area of photocatalyst, dissolve oxygen in the reaction medium, nature of the photocatalyst, nature of the substrate, doping of metal ions and non metal and structure of photocatalyst and substrate. The photodegradation of organic compound have been studied by the several scientists and conclude the optimum conditions for the photodegradation of organic compound.

The photodegradation of organic compounds was found maximum at low concentration of organic substrate with optimum amount of photocatalyst. The pH of the solution is also affect the photodegradation of organic substrate. The Titania show the maximum adsorption at low pH hence the photodegradation also found maximum at low pH. The surface area is the crucial factor for the photodegradation of organic substrate. If we are increasing the surface area of photocatalyst, the photodegradation of organic substrate increase. This is because that the number of active site increased with increasing the surface area. The amount of photocatalyst is the primary requirement of any photocatalysis process. The amount of photocatalyst should be optimum, if we take the high amount of photocatalyst the photodegradation is decreased, if we take the low amount of photocatalyst the photodegradation also decreased. The doping of metal ions and non-metal ions affect the photodegradation of organic substrate. Therefore, we have to use the metal ions which increased the positive charge on the surface of photocatalyst. The temperature and irradiation of light also affect the photodegradation of organic substrate. For the maximum photodegradation, the photocatalytic reaction should perform at room temperature not greater than 80°C. The light of irradiation should be used which exit the electron easily from valence band to conduction band or equal to band gap energy.

**Keywords:** photodegradation, photocatalyst, surface area, nanocomposites, titania, metal doped titania

Volume I Issue 3 - 2017

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**Received:** August 10, 2017 | **Published:** November 08, 2017

## Introduction

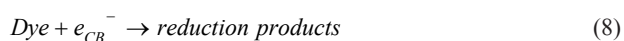
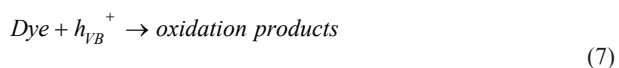
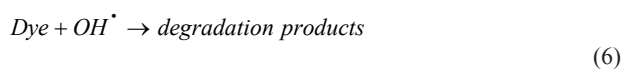
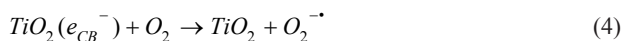
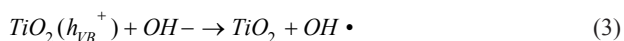
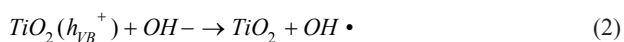
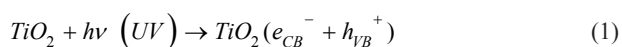
The industrial waste water containing the various dyes is responsible for water pollution, due to their carcinogenic behaviour. Lots of investigations reported that 10-12% of dyes are used every year in textile industries such as Rose Bengal, Victoria blue, Thymol blue, Caramine, Indigo Red, Red 120, Rhodamine B, Methylene Blue, Eriochrome Black-T (EBT)<sup>1-4</sup> of which are major portion (20%) lost during synthesis and processing operations, which enter into water through effluents. The wastewater released from industries contains highly hazardous and coloured pigments which causes serious ill effect on aquatic life and also human beings.<sup>5-8</sup>

In the past, several physical techniques like photo degradation, **coagulation**, **floculation**, reverse osmosis, adsorption on the activated carbon, ion exchange method ultra-filtration and chemical methods like photosensitized oxidation, adsorption,<sup>9-12</sup> have been used to reduce the toxic dye effluents from waste water though methods are fairly effective in removing pollutants. However the main drawback of these techniques is formation of secondary waste product which cannot be treated again and dumped as such.<sup>13,14</sup>

The photo catalysis is the very advanced oxidation processes (AOP) which is used for the photo degradation of toxic compounds. It is also used for the purification of water.<sup>15-17</sup> Basically, the photo catalysis is divided into two categories (1) heterogeneous catalysis and (2) homogenous catalysis. Heterogeneous catalysis has been successfully employed for the degradation of various families of hazardous materials. The photodegradation has lot of advantages over traditional wastewater treatment techniques such as chemical oxidation,<sup>18</sup> activated carbon adsorption,<sup>19</sup> biological treatment,<sup>20</sup> etc. Activated carbon adsorption method involves the phase transfer of pollutants without decomposition and thus creates another pollution problem. Chemical oxidation method is unable to remove all organic substances and it is suitable for the removal of pollutants at high concentrations. The biological treatments are very slow, dispose large amount of sludge and required strict control of proper pH and temperature.<sup>21</sup> In this regard photo catalytic processes have advantages for the removal of pollutants even at low concentration for industrial waste water.<sup>22</sup> Moreover in photo oxidation, complete oxidation of organic pollutants take place within few hours, even at ppb level, without formation of secondary hazardous products using

highly active and cheap catalysts which can be used in specially design reactor systems.<sup>23</sup>

Titanium dioxide is a widely accepted photocatalyst due to its high oxidation efficiency, non-toxicity, high photo stability, chemical inertness and environmental friendly nature.<sup>24,25</sup> It is a wide band gap (~3.2eV) semiconductor and mineralizes a large range of organic pollutants such as herbicides, dyes, pesticides, phenolic compounds, tetracycline, sulfamethazine, etc under UV irradiation.<sup>26</sup> In photocatalytic processes the photon with energy, equal to or greater than the energy gap (3.2eV) is absorbed by TiO<sub>2</sub> particles, therefore holes (h<sup>+</sup>) are created in the valence band and electrons (e<sup>-</sup>) in the conduction band, which diffuses to the surface of TiO<sub>2</sub> and participates in redox reactions of the adsorbed substrates.<sup>27</sup> In brief, photo generation of radical species in TiO<sub>2</sub>/UV system can be described as follows.<sup>28,29</sup>



Where hv is the photon energy which is required to excite the semiconductor electron from the valence band (VB) to conduction band (CB).

### Factors influencing the photocatalytic degradation

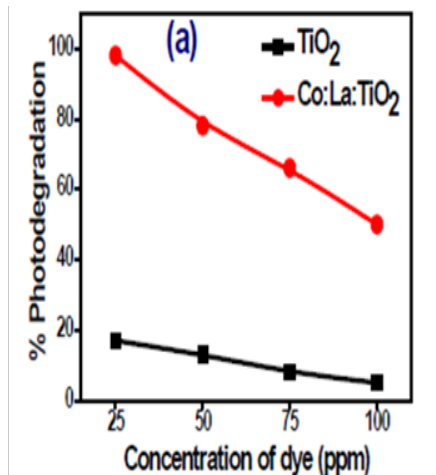
The oxidation rates and efficiency of the photocatalytic system are highly dependent on a number of operational parameters that govern the photodegradation of the organic molecule.<sup>30-34</sup> Several study have been reported the significance of operational parameter.

#### Effect of dye concentration

The photocatalysis is depends on the adsorption of dyes on the surface of photocatalyst. In the photocatalysis process, only the amount of dye adsorbed on the surface of photocatalyst contributes and not the one in the bulk of the solution. The adsorption of dye depends on the initial concentration of dye. The initial concentration of dye in a given photocatalytic reaction is an important factor which needs to be taken into account. Generally speaking the percentage degradation decreases with increasing amount of dye concentration, while keeping a fixed amount of catalyst.<sup>35</sup>

This can be rationalized on the basis that as dye concentration

increases, more organic substances are adsorbed on the surface of TiO<sub>2</sub>, whereas less number of photons are available to reach the catalyst surface and therefore less •OH are formed, thus resulting in less degradation percentage. Figure 1 is showing the photodegradation of methyl orange with different photocatalyst. The photodegradation of dyes are increase with decreasing the concentration of dyes with the photocatalyst.<sup>36</sup>



**Figure 1** Effect of concentration on photodegradation Methyl Orange with TiO<sub>2</sub> and Co:La:TiO<sub>2</sub>.<sup>36</sup>

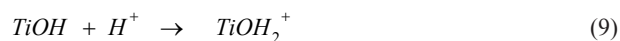
#### Effect of catalyst amount

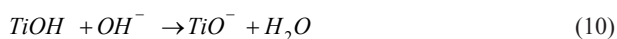
Degradation of dye is affected by the amount of the photocatalyst. The photodegradation of dye increases with increasing catalyst amount, which is the feature of heterogeneous photocatalysis. The increase in catalyst amount actually increases the number of active sites on the photocatalyst surface thus causing an increase in the formation of number of •OH radicals which can take part in actual discoloration of dye solution. Beyond a certain limit of catalyst amount, the solution becomes turbid and thus blocks UV radiation for the reaction to proceed and therefore percentage degradation starts decreasing.<sup>37</sup>

Figure 2 showing the photodegradation of methyl orange dye at different amount. The percent photodegradation is increase with increasing the amount of photocatalyst.<sup>36</sup> Figure 3 is also showing the effect of photocatalyst amount on the photodegradation of Eriochrome black T (EBT). The Photodegradation of EBT is also increase with amount of photocatalyst.<sup>38</sup> Figure 4 showing the photodegradation of acetic acid in presence of Titania and Co, Ni modified Titania at different amount. The photodegradation of acetic acid was increased with increase of photocatalyst amount.<sup>39</sup>

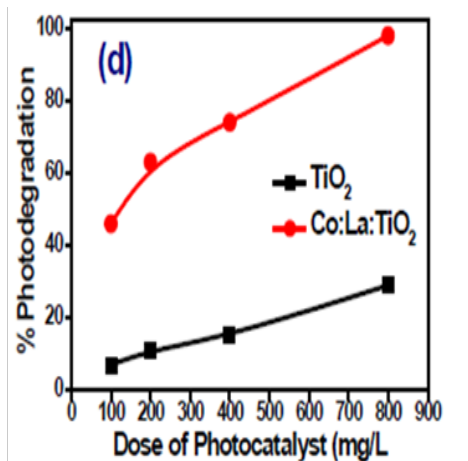
#### Effect of pH

The Photodegradation of dyes are affected by the pH of the solution. The variation of solution pH changes the surface charge of TiO<sub>2</sub> particles and shifts the potentials of catalytic reactions. As a result, the adsorption of dye on the surface is altered thereby causing a change in the reaction rate. Under acidic or alkaline condition the surface of Titania can be protonated or deprotonated respectively according to the following reactions:<sup>40</sup>





Thus titania surface will remain positively charged in acidic medium and negatively charged in alkaline medium. Titanium dioxide is reported to have higher oxidizing activity at lower pH, but excess  $\text{H}^+$  can decrease reaction rate.  $\text{TiO}_2$  behaves as a strong Lewis acid due to the surface positive charge. In other words, the anionic dye acts as a strong Lewis base and can easily adsorb on the positively charged catalyst surface. This favours the adsorption of the dye under acidic conditions, while in the alkaline conditions this complexation process is not favoured presumably because of competitive adsorption by hydroxyl groups and the dye molecule in addition to the Coulombic repulsion due to the negatively charged catalyst with the dye molecule.<sup>41</sup> The extent of dye adsorption depends on the initial dye concentration, nature of the dye, surface area of photocatalyst and pH of the solution. The pH determines the surface charge of the photocatalyst. Adsorption of the dye is minimum when the pH of the solution is at the isoelectric point (point of zero charge).<sup>42</sup> The surface of the photocatalyst is positively charged below isoelectric point and carries a negative charge above it. Figure 5–7 showing the effect of pH on the photodegradation of methyl red,<sup>43</sup> EBT<sup>38</sup> and Oxalic Acid.<sup>44</sup> The Photodegradation of Methyl Red is found maximum at 5pH because the methyl Red is the cationic dye which adsorb easily on the surface of Titania (Figure 5). The photodegradation of EBT is found maximum at 2pH because it is anionic dye which adsorb maximum at 2pH. Hence maximum photo-degradation occurs at 2pH (Figure 6). Figure 7 showing the effect of pH on photodegradation of oxalic acid at two temperatures with Cu-TiO<sub>2</sub>. The photodegradation of Oxalic acid was found maximum at 4pH. This is because that in acetic medium the adsorption of Oxalic acid is found maximum. Therefore, the photodegradation of oxalic acid was found maximum at 4pH.

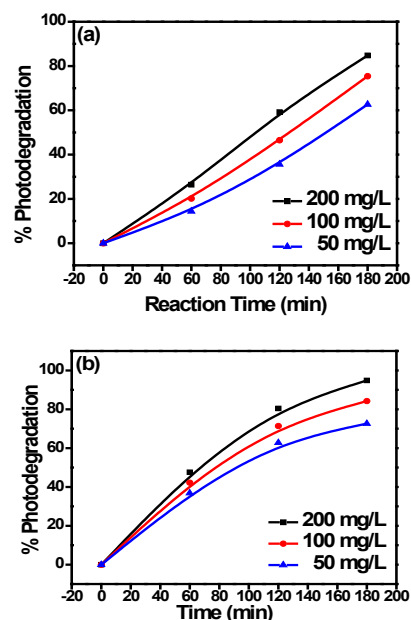


**Figure 2** Effect of photocatalyst amount on photodegradation of Methyl Orange with (a)  $\text{TiO}_2$  (b) Co:La:TiO<sub>2</sub>.<sup>36</sup>

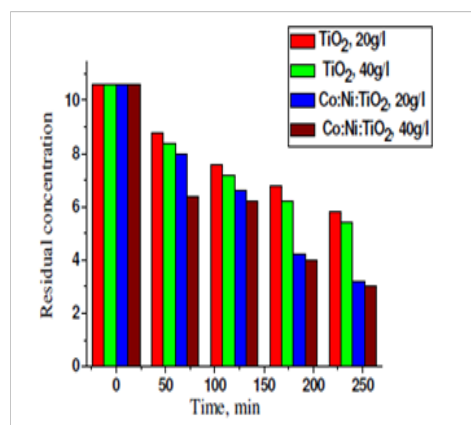
### Size and structure of the photocatalyst

Surface morphology such as particle size and agglomerate size, is an important factor to be considered in photocatalytic degradation process because there is a direct relationship between organic compounds and surface coverage of the photocatalyst.<sup>45</sup> The number of photon striking the photocatalyst controls the rate of reaction which signifies that the reaction takes place only in the absorbed phase of the photocatalyst.<sup>46</sup> A number of different forms of  $\text{TiO}_2$  have been synthesized to achieve the desired characteristics of the

photocatalyst.<sup>47</sup> Figure 8 showing the photodegradation of methyl red dyes with Titania and Cu-TiO<sub>2</sub> nanocomposites. The maximum photodegradation is found in case of Cu modified Titania and whereas in case of Titania less photodegradation occurs.<sup>43</sup> This is because that the surface of Titania different from Cu modified Titania. The Cu modified Titania has rough and zigzag surface which is help in the adsorption of organic molecule on the catalyst surface. So that Cu modified Titania shows good photocatalytic activity than pure Titania.



**Figure 3** Effect of photocatalyst amount on photodegradation of EBT with (a)  $\text{TiO}_2$  (b) Ni:TiO<sub>2</sub>.<sup>38</sup>



**Figure 4** Effect of photocatalyst amount on photodegradation of acetic acid with  $\text{TiO}_2$  and Co:Ni:TiO<sub>2</sub>.<sup>39</sup>

Figure 9 shows the effect of different structures on the photodegradation of EBT. The nanoparticle of Titania which is P25-TiO<sub>2</sub> is the superior photocatalyst than the Titania nanotubes and nanorods. This is because that the nanoparticles have higher surface area than the nanotubes and nanorods. The nanoparticles show the 86-94% degradation of EBT.

### Surface area

Surface morphology of  $\text{TiO}_2$  is a crucial factor in its use as photocatalyst, as all the chemical events take place at the surface,

its enlargement has been attempted, usually by using very fine particles, either suspended in solvents or made into a porous film. Nanostructured materials with the crystallite/grain size below 20nm are of great research interest mainly due to the fact that their physical properties may be markedly different from the bulk counter parts. This has also opened up avenues for their applications as photocatalyst in numerous areas.<sup>48</sup> Figure 10 shows the effect of surface area on photodegradation of rose Bengal dye in presence of PANi/Graphene nanocomposites. The surface area of photocatalyst increases the photodegradation of rose Bengal also increase. This is because that the high surface area contains the greater no of active site than the low surface area materials.

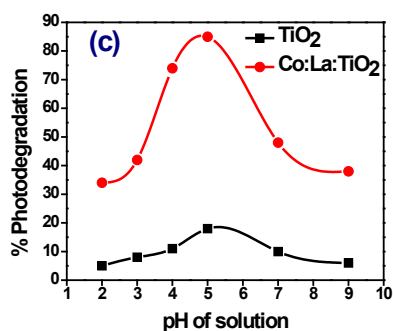


Figure 5 Effect of pH of solution on photodegradation of methyl red with TiO<sub>2</sub> and Co:La:TiO<sub>2</sub>.<sup>43</sup>

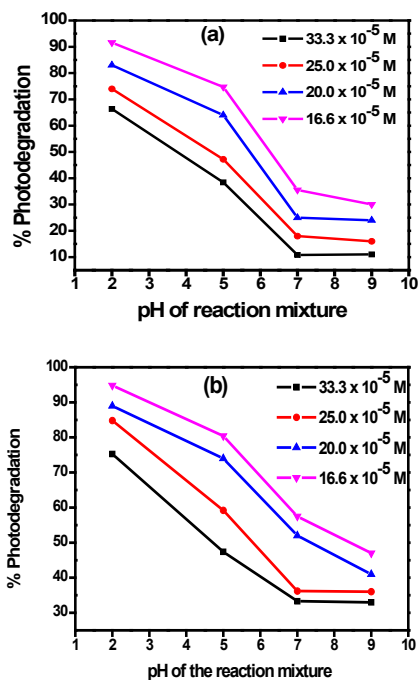


Figure 6 Effect of pH on photodegradation of EBT at various concentrations (a) TiO<sub>2</sub> (b) Ni:TiO<sub>2</sub>.<sup>38</sup>

### Reaction temperature

An increase in reaction temperature generally results in increased photocatalytic activity however reaction temperature >80°C promotes the recombination of charge carriers and disfavours the adsorption of organic compounds on the Titania surface.<sup>49</sup> A reaction temperature

below 80°C favours the adsorption whereas further reduction of reaction temperature to 0°C results in an increase in the apparent activation energy.<sup>50</sup> Therefore temperature range between 20-80°C has been regarded as the desired temperature for effective photo mineralization of organic content.<sup>51</sup>

Figures 11 and 12 showing the effect of temperature on the photo-degradation of methyl red and methyl orange in presence of TiO<sub>2</sub> and La and Co modified Titania. At the high temperature the photodegradation was found maximum whereas at low temperature the photodegradation of dyes was found low efficiency of photodegradation. This is due to the interaction of dyes molecule increase with the photocatalyst surface. Hence the adsorption increases which favour the photodegradation of dyes.<sup>36,43</sup>

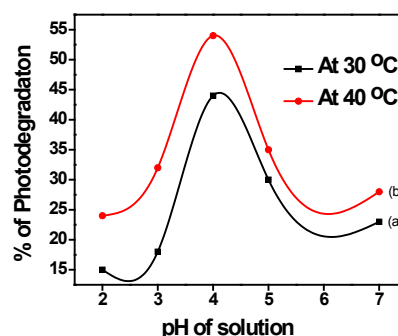


Figure 7 The effect of pH on the photodegradation of oxalic acid at different temperature with Cu-TiO<sub>2</sub>.<sup>44</sup>

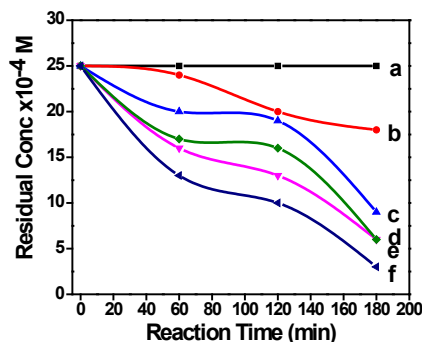


Figure 8 Photodegradation of methyl red at initial concentration  $25 \times 10^{-4} \text{M}$  (a) without photocatalyst at 30°C (b) without photocatalyst at 40°C (c) TiO<sub>2</sub> as photocatalyst at 30°C (d) TiO<sub>2</sub> as photocatalyst at 40°C (e) Cu-TiO<sub>2</sub> as photocatalyst at 30°C (f) Cu-TiO<sub>2</sub> as photocatalyst at 40°C.

### Concentration and nature of pollutants

The rate of photocatalytic degradation of certain pollutant depends on its nature, concentration and other existing compounds in water matrix. A number of studies have reported the dependency of the TiO<sub>2</sub> reaction rate on the concentration of contaminants in water.<sup>52</sup> High concentration of pollutants in water saturates the TiO<sub>2</sub> surface and hence reduces the photonic efficiency and deactivation of the photocatalyst.<sup>53</sup> In addition to the concentration of pollutants, the chemical structure of the target compound also influences the degradation performance of the photocatalytic reactor. For example, 4-chlorophenol requires prolonged irradiation time due to its transformation to intermediates compared with oxalic acid that transforms directly to carbon dioxide and water, i.e., complete mineralization.<sup>54</sup> Furthermore if the nature of the target water contaminants is such that they adhere effectively

to the photocatalyst surface the process would be more effective in removing such compounds from the solution. The photocatalytic degradation of aromatics is highly dependent on the substituent group.<sup>55</sup> The organic substrates with electron withdrawing nature (benzoic acid, nitrobenzene) strongly adhere to the photocatalyst and therefore are more susceptible to direct oxidation compared with the electron donating groups.<sup>56</sup>

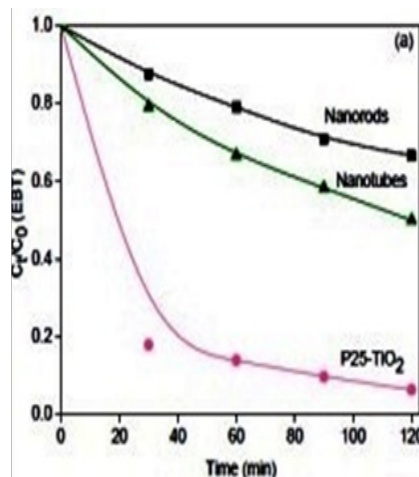


Figure 9 Effect of structure of the photocatalyst on the photodegradation of EBT.<sup>32</sup>

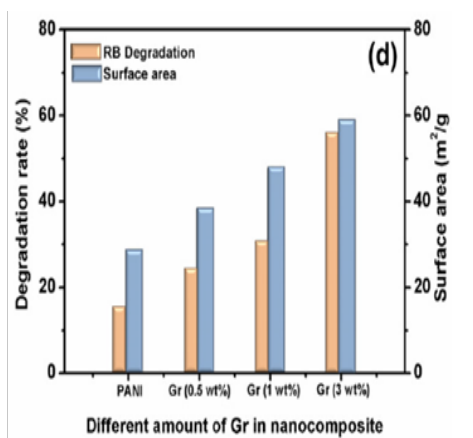


Figure 10 The Effect of surface area on the photodegradation of Rose Bengal dye.<sup>48</sup>

### Inorganic ions

Various inorganic ions such as magnesium, iron, zinc, copper, bicarbonate, phosphate, nitrate, sulphate and chloride present in wastewater can affect the photocatalytic degradation rate of the organic pollutants because they can be adsorbed onto the surface of TiO<sub>2</sub>.<sup>57</sup> Photocatalytic deactivation has been reported whether photocatalyst is used in slurry or fixed-bed configuration which is related to the strong inhibition from the inorganic ions on the surface of the TiO<sub>2</sub>.<sup>58</sup> A number of studies have been conducted on the effect of inorganic ions (anions and cations) on TiO<sub>2</sub> photocatalytic degradation.<sup>59</sup> Some of the cations such as copper, iron and phosphate have been reported to decrease the photodegradation efficiency if they are present at certain concentrations whereas calcium, magnesium and zinc have little effect on the photodegradation of organic compounds which is associated to the fact that these cations have are at their maximum

oxidation states that results in their inability to have any inhibitory effect on the degradation process.<sup>60</sup>

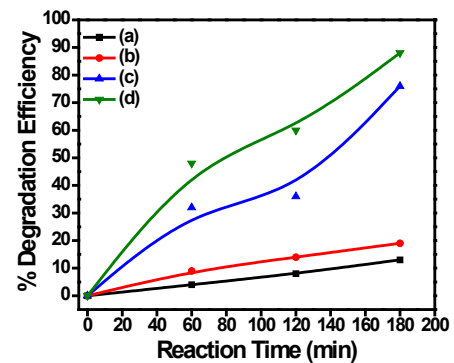


Figure 11 Photodegradation of methyl red at initial concentration 50 ppm (a) TiO<sub>2</sub> at 30°C (b) TiO<sub>2</sub> at 40°C (c) Co:La:TiO<sub>2</sub> at 30°C (d) Co:La:TiO<sub>2</sub> at 40°C.<sup>43</sup>

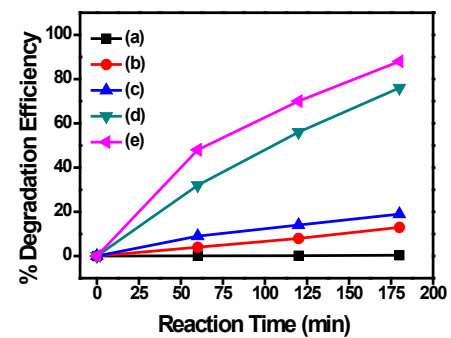


Figure 12 The Effect of surface area on the photodegradation of rose bengal dye.<sup>48</sup>

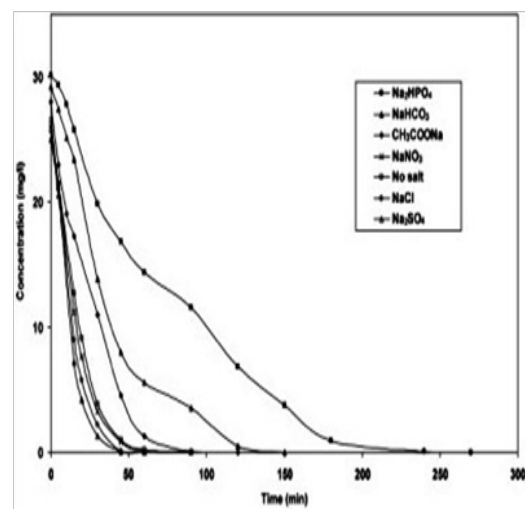
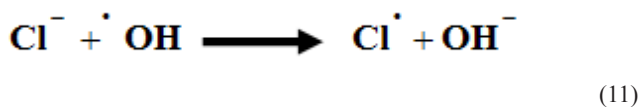


Figure 13 Effect of anions on the photodegradation of RY 145 (C<sub>0</sub>=30mg/l).<sup>67</sup>

The inorganic anions such as nitrate, chlorides, carbonates and sulphates are also known to inhibit the surface activity of the photocatalyst. The presence of salts diminishes the colloidal stability, increases mass transfer and reduces the surface contact between the pollutant and the photocatalyst.<sup>61</sup> Other than fouling of the TiO<sub>2</sub> surface certain anions such as chlorides, carbonates, phosphate and sulphates also scavenge both the hole and the hydroxyl radicals.<sup>62</sup> The

mechanism of whole and radical scavenging by chloride has been proposed by Matthews and McEnvoy<sup>63</sup> as follows.



The inhibitory effect of chloride ions occurs through preferential adsorption displacement mechanism which results in reducing the number of OH<sup>-</sup> available on the photocatalyst surface.<sup>64</sup> The fouling of photocatalytic surface can be reduced by pre-treatment of water such as with ion exchange resins which have been reported to reduce the fouling and so the cost of treatment.<sup>65</sup> Similarly the fouling induced by sulphates and phosphates has been reported to be displaced by NaOH, KOH and NaHCO<sub>3</sub>.<sup>66</sup> However, most of studies conducted on the effect of inorganic ions are based on the model compounds and therefore do not necessarily represent their effect in real water matrix where several ions exist. More work concentrating on the effect of complex mixtures of inorganic ions is thus required. Figure 13 showing the effect of different ions on the photodegradation of RY 145 dye at initial concentration of 30mg/L.

### Effect of light intensity and irradiation time

Both light intensity and time of irradiation affect the dye degradation.<sup>68</sup> It has been shown that at low light intensities (0-20mW/cm<sup>2</sup>), the rate would increase linearly with increasing light intensity (first order), whereas at intermediate light intensities (25mW/cm<sup>2</sup>) the rate would depend on the square root of the light intensity.<sup>69</sup> At high light intensities the rate is independent of light intensity, because at low light intensity reactions involving electron-hole formation are predominant and electron-hole recombination is negligible. On the other hand, when light intensity is increased, the electron-hole pair separation competes with recombination, thereby causing lower effect on the reaction rate.

Figure 14 shows the effect of light intensity on the photodegradation of 2,4-D. The photodegradation of 2,4-D is increased with increasing the light intensity from 100 to 600 Lux. In many literature studies, it has been shown that the dye decolorization initially increases as the light intensity is increased.<sup>70</sup>

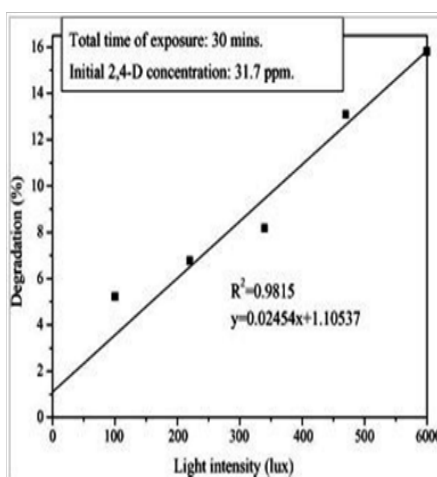


Figure 14 Effect of light intensity on 2,4-D photodegradation.<sup>70</sup>

The reaction rate decreases with irradiation time as it follows the pseudo first-order kinetics and additionally a competition for degradation may occur between the reactant and the intermediate products. The slow kinetics of dye degradation after certain time limit is mainly attributed to the difficulty in the reaction of short chain aliphatics with  $\cdot\text{OH}$  radicals, and the short lifetime of photocatalyst because of active sites deactivation by strong by-products deposition.<sup>71</sup> Figure 15 shows the effect of irradiation time on the photodegradation of methyl orange with Titania and Co and La modified Titania. The photodegradation of methyl orange was found highest at 3 hour irradiation of visible light on the photocatalytic reaction. Hence this is confirm that the photodegradation increase with increase of time of irradiation on the photocatalytic reaction.

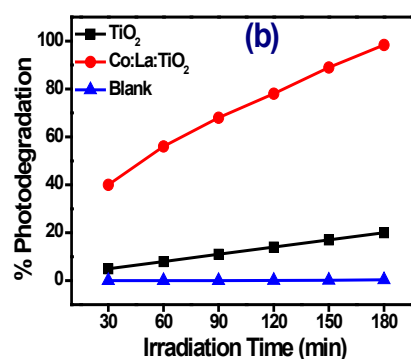


Figure 15 Effect of irradiation time on photodegradation of methyl orange.<sup>36</sup>

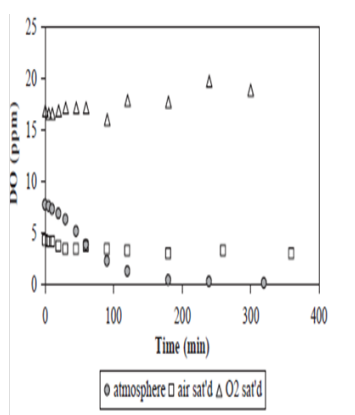
### Dissolved oxygen

Oxygen dissolved in solution is commonly employed as an electron acceptor in photocatalysis reaction to assure sufficient electron scavengers present to trap the excited conduction band electron from recombination.<sup>72</sup> The oxygen does not affect the adsorption on the TiO<sub>2</sub> catalyst surface as the reduction reaction takes place at a different location from where oxidation occurs. Dissolved oxygen involves in the stabilization of radical intermediates, mineralization and direct photocatalytic reactions. Its presence is also known to induce the cleavage mechanism for aromatic rings in organic pollutants that are present in the water matrices.<sup>73</sup> Figure 16 shows the Effect of dissolved oxygen concentration on photocatalytic degradation, at initial dye concentration=250ppm, TiO<sub>2</sub> loading=1g/L, flow rate=1L/min. The photodegradation of dye increased with increasing the DO. This is due to the formation of Oxygen free radical which increases the photodegradation.

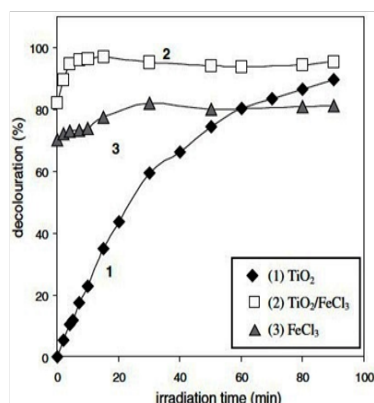
### Effect of dopants on dye degradation

Heterogeneous photocatalysis involving titanium dioxide (TiO<sub>2</sub>) appears to be the most promising technology for organic dyes degradation. However one of the major problems in using TiO<sub>2</sub> as a catalyst is the low photo-quantum efficiency which arises from the fast recombination of photogenerated electrons and holes. Moreover, TiO<sub>2</sub> is inactive under visible light due to its wide band gap (3.03eV for rutile and 3.18 for anatase form). This inherently causes the inability to make use of the vast potential of solar photocatalysis. Various techniques have been employed to make TiO<sub>2</sub> absorb photons of lower energy as well. These techniques include surface modification via organic materials and semiconductor coupling, band gap modification by creating oxygen vacancies and oxygen sub-stoichiometry, by nonmetals including co-doping of nonmetals and metal doping.

Dopants, such as transitional metals have been added to the TiO<sub>2</sub> catalyst to improve its response and also reduce the recombination of photogenerated electrons and photogenerated holes.<sup>74,75</sup> The main objective of doping is to induce a bathochromic shift, i.e., a decrease of the band gap or introduction of intra-band gap states, which results in more visible light absorption. The effect of metal ion dopants on the photocatalytic activity is a complex problem. The total induced alteration of the photocatalytic activity is made up from the sum of changes which occur in the light absorption capability of the TiO<sub>2</sub> photocatalyst, adsorption capacity of the substrate molecules at the catalyst's surface and interfacial charge transfer rate. Using suitable transitional metals as dopants improves the performances of TiO<sub>2</sub>.<sup>76,77</sup> This results in the overlap of the conduction band due to Ti (3d) with d levels of the transition metals causing red shift of the band edge of TiO<sub>2</sub>. It can thus also allow the light absorption to be widened into the visible region to various extents, depending on the type of the dopant and its concentration. Therefore the photocatalysis on modified TiO<sub>2</sub> can be promoted using visible light. Various transition metals, such as Ni<sup>2+</sup>, Zn<sup>2+</sup>, Cr<sup>3+</sup> and Fe<sup>3+</sup> can be easily incorporated into the crystal lattice of TiO<sub>2</sub> because of their similar ionic radii (The ionic radii of Ni<sup>2+</sup>=0.72Å, Zn<sup>2+</sup>=0.74Å, Cr<sup>3+</sup>=0.76Å and Fe<sup>3+</sup>=0.69 Å are quite similar to that of host Ti<sup>4+</sup>=0.75Å ions).<sup>78–81</sup> The degradation of dyes is usually faster in mixed systems as compared to single systems, because the oxidation of dyes consumes photo-excited holes promptly and efficiently, thus attenuating electron–hole recombination.<sup>82</sup> Figure 17 showing the Effect of Cl ion on the photodegradation of acid orange dye. The photodegradation increased with the impregnation of Cl ion in Titania.<sup>83</sup>



**Figure 16** Effect of dissolved oxygen concentration on photocatalytic degradation, initial dye concentration=250ppm, TiO<sub>2</sub> loading=1g/L, flow rate=1L/min.<sup>74</sup>



**Figure 17** Comparison of decolouration course of acid orange 7 solutions (C0<sup>1/4</sup> 0:1mM) in the presence of TiO<sub>2</sub> (2.5g/l), FeCl<sub>3</sub> (1.0mM) and TiO<sub>2</sub> (2g/l)/FeCl<sub>3</sub> (1.0mM).<sup>83</sup>

## Acknowledgements

None.

## Conflict of interest

The author declares no conflict of interest.

## References

- Konstantinou IK, Albanis TA. TiO<sub>2</sub>-assisted photocatalytic degradation of azo dyes in aqueous solution: kinetic and mechanistic investigations: a review. *Applied Catalysis B Environmental*. 2004;49(1):1–14.
- Deng F, Min L, Luo X, et al. Visible–light photocatalytic degradation performances and thermal stability due to the synergetic effect of TiO<sub>2</sub> with conductive copolymers of polyaniline and polypyrrole. *Nanoscale*. 2013;5(18):8703–8710.
- Xu S, Zhu Y, Jiang L, et al. Visible light induced photocatalytic degradation of methyl orange by polythiophene/TiO<sub>2</sub> composite particles. *Water, Air & Soil Pollution*. 2010;213(1–4):151–159.
- Wang D, Wang Y, Li X, et al. Sunlight photocatalytic activity of polypyrrole–TiO<sub>2</sub> nanocomposites prepared by ‘in situ’ method. *Catalysis Communications*. 2008;9(6):1162–1166.
- Park Y, Lee S, Kang SO, et al. Organic dye–sensitized TiO<sub>2</sub> for the redox conversion of water pollutants under visible light. *Chemical Communications*. 2010;46(14):2477–2479.
- Hoffmann MR, Martin ST, Choi W, et al. Environmental applications of semiconductor photocatalysis. *Chemical Reviews*. 1995;95(1):69–96.
- Linsebigler AL, Lu GQ, Yates JT. Photocatalysis on TiO<sub>2</sub> surfaces: principles, mechanisms and selected results. *Chem Rev*. 1995;95(3):735–758.
- Stewart G, Fox MA. The effect of dark recovery time on the photoefficiency of heterogeneous photocatalysis by TiO<sub>2</sub> suspended in non-aqueous media. *Res Chem Intermediates*. 1995;21:933–938.
- Serpone N. A decade of heterogeneous photocatalysis in our laboratory: pure and applied studies in energy production and environmental detoxification. *Research on Chemical Intermediates*. 1994;20(9):953–992.
- Zamaraev KI, Khramov MI, Parmon VN. Possible impact of heterogeneous photocatalysis on the global chemistry of the earth's atmosphere. *Cat Rev Sci Eng*. 1994;36(4):617–644.
- Kamat PV. Photochemistry on nonreactive and reactive (semiconductor) surfaces. *Chem Rev*. 1993;93(1):267–300.
- Liang HC, Li XZ. Visible–induced photocatalytic reactivity of polymer–sensitized titania nanotube films. *Applied Catalysis B: Environmental*. 2009;86(1–2):8–17.
- Ferreira C, Domenech S, Lacaze P. Synthesis and characterization of polypyrrole/TiO<sub>2</sub> composites on mild steel. *J of Applied Electrochemistry*. 2001;31(1):49–56.
- Saqib M, Muneer M. TiO<sub>2</sub>-mediated photocatalytic degradation of a triphenylmethane dye (gentian violet), in aqueous suspensions. *Dyes Pigments*. 2003;56(1):37–49.
- Li X, Liu G, Zhao J. Two competitive primary processes in the photo-degradation of cationic triaryl methane dyes under visible irradiation in TiO<sub>2</sub> dispersions. *New J Chem*. 1999;23:1193–1196.
- Deng F, Li Y, Luo X, et al. Preparation of conductive polypyrrole/TiO<sub>2</sub> nanocomposite via surface molecular imprinting technique and its photocatalytic activity under simulated solar light irradiation. *Colloids and Surfaces A: Physicochem Eng Aspects*. 2012;395:183–189.
- Chen CC, Lu CS. Photocatalytic degradation of Basic Violet 4: degradation efficiency, product distribution, and mechanisms. *J Phys Chem C*. 2007;111(37):13922–13932.

18. Okano M, Itoh K, Fujishima A, et al. Photoelectrochemical polymerization of pyrrole on TiO<sub>2</sub> and its application to conducting pattern generation. *Journal of the Electrochemical Society*. 1987;134(4):837–841.
19. Zeltner WA, Hill CG, Anderson MA. Supported titania for photodegradation. *Chem Tech*. 1993;23:21.
20. Mills A, Davies RH, Worsley D. Water purification by semiconductor photocatalysis. *Chem Soc Rev*. 1993;22(6):417.
21. *Photocatalytic Purification and Treatment of Water and Air*. In: Ollis DF, Al-Ekabi H, editors. Netherlands: Elsevier; 1993.
22. *Photocatalysis: Fundamentals and Applications*. In: Serpone N, Pelizzetti E, editors. USA: Wiley; 1989.
23. Jia H, Zheng Z, Zha H, et al. Aqueous sol–gel synthesis and growth mechanism of single crystalline TiO<sub>2</sub> nanorods with high photocatalytic activity. *Materials Research Bulletin*. 2009;44(6):1312–1316.
24. Cai H, Mu W, Liu W, et al. Sol–gel synthesis highly porous titanium dioxide microspheres with cellulose nano fibrils–based aerogel templates. *Inorganic Chemistry Communications*. 2015;51:71–74.
25. Jongprateep O, Puranasamriddhi R, Palomas J. Nanoparticulate titanium dioxide synthesized by sol–gel and solution combustion techniques. *Ceramics International*. 2015;41(1):S169–S173.
26. Lee H, Park YK, Kim SJ, et al. Titanium dioxide modification with cobalt oxide nanoparticles for photocatalysis. *Journal of Industrial and Engineering Chemistry*. 2015;32:259–263.
27. Makwana MN, Christopher JT, Robert IG, et al. Pilot plant scale continuous hydrothermal synthesis of nano–titania; effect of size on photocatalytic activity. *Materials Science in Semiconductor Processing*. 2016;42(1):131–137.
28. Akpan UG, Hameed BH. Parameters affecting the photocatalytic degradation of dyes using TiO<sub>2</sub>–based photocatalysts: A review. *Journal of Hazardous Materials*. 2009;170(2–3):520–529.
29. Fujishima A, Zhang X, Tryk DA. TiO<sub>2</sub> photocatalysis and related surface phenomena. *Surface Science Reports*. 2008;63(12):515–582.
30. Ng CM, Chen PC, Manickam S. Hydrothermal crystallization of titania on silver nucleation sites for the synthesis of visible light nano–photocatalysts–Enhanced photoactivity using Rhodamine 6G. *Applied Catalysis A: General*. 2012;433–434:75–80.
31. Gnanaprakasam A, Sivakumar VM, Sivayogavalli PL, et al. Characterization of TiO<sub>2</sub> and ZnO nanoparticles and their applications in photocatalytic degradation of azodyes. *Ecotoxicology and Environmental Safety*. 2015;121:121–125.
32. Kaur S, Singh V. Visible light induced sonophotocatalytic degradation of Reactive Red dye 198 using dye sensitized TiO<sub>2</sub>. *Ultrasonics Sonochemistry*. 2007;14(5):531–537.
33. Kumar A, Pandey G. Synthesis of La:Co:TiO<sub>2</sub> Nanocomposite and Photocatalytic Degradation of Tartaric Acid in Water at Various Parameters. *American Journal of Nano Research and Applications*. 2017;5(4):40–48.
34. Grzechulska J, Morawski AW. Photocatalytic decomposition of azo–dye acid black I in water over modified titanium dioxide. *Applied Catalysis B: Environmental*. 2002;36(1):45–51.
35. Reza KM, Kurny ASW, Gulshan F. Parameters affecting the photocatalytic degradation of dyes using TiO<sub>2</sub>: a review. *Applied Water Science*. 2017;7(4):1569–1578.
36. Kumar A, Pandey G. Photodegradation of Methyl Orange in Aqueous Solution by the Visible Light Active Co:La:TiO<sub>2</sub> Nanocomposite. *Chem Sci J*. 2017;8:164.
37. Coleman HM, Vimonse V, Leslie G, et al. Degradation of 1,4–dioxane in water using TiO<sub>2</sub> based photocatalytic and H<sub>2</sub>O<sub>2</sub>/UV processes. *J Hazard Mater*. 2007;146(3):496–501.
38. Kumar A, Pandey G. Photocatalytic degradation of Eriochrome Black–T by the Ni:TiO<sub>2</sub> nanocomposites. *Desalination and Water Treatment*. 2017;71:406–419.
39. Kumar A, Pandey G. Photocatalytic Activity of Co:TiO<sub>2</sub> Nanocomposites and their Application in Photodegradation of Acetic Acid. *Chemical Science Transactions*. 2017;6(3):385–392.
40. Davis RJ, Gainer JL, Neal GO, et al. Photocatalytic decolorization of wastewater dyes. *Water Environ Res*. 1994;66(1):50–53.
41. Mozia S, Morawski AW, Toyoda M, et al. Application of anatase–phase TiO<sub>2</sub> for decomposition of azo dye in a photocatalytic membrane reactor. *Desalination*. 2009;241(1–3):97–105.
42. Cavigli L, Bogani F, Vinattieri A, et al. Volume versus surface–mediated recombination in anatase TiO<sub>2</sub> nanoparticles. *Journal of Applied Physics*. 2009;106(5):053516.
43. Kumar A, Hitkari G, Gautam M, et al. Synthesis, Characterization and Application of Cu–TiO<sub>2</sub> Nanocomposites in Photodegradation of Methyl Red (MR). *International Advanced Research Journal in Science Engineering and Technology*. 2015;2(12):50–55.
44. Kumar A, Hitkari G, Gautam M, et al. Synthesis of Ni–TiO<sub>2</sub> Nanocomposites and Photocatalytic Degradation of Oxalic Acid in Waste Water. *International Journal of Innovative Research in Science, Engineering and Technology*. 2015;4(12):12721–12731.
45. Kormann C, Bahnemann DW, Hoffman MR. Photocatalytic production of H<sub>2</sub>O<sub>2</sub> and organic peroxides in aqueous suspensions of TiO<sub>2</sub>, ZnO and desert sand. *Environmental Science & Technology*. 1988;22(7):798–806.
46. Zhu J, Deng Z, Chen F, et al. Hydrothermal doping method for preparation of Cr<sup>3+</sup>–TiO<sub>2</sub> photocatalysts with concentration gradient distribution of Cr<sup>3+</sup>. *Applied Catalysis B: Environmental*. 2006;62(3–4):329–335.
47. Nakata K, Fujishima A. TiO<sub>2</sub> photocatalysis: Design and applications. *Journal of Photochemistry and Photobiology C: Photochemistry Reviews*. 2012;13(3):169–189.
48. Ameen S, Seo HK, Akhtar MS, et al. Novel graphene/polyaniline nanocomposites and its photocatalytic activity toward the degradation of rose Bengal dye. *Chemical Engineering Journal*. 2012;210(1):220–228.
49. Hashimoto K, Irie H, Fujishima A. TiO<sub>2</sub> photocatalysis: a historical overview and future prospects. *Japanese Journal of Applied Physics*. 2005;44(1):8269–8285.
50. Peral J, Domenech X, Ollis DF. Heterogeneous photocatalysis for purification, decontamination and deodorization of air. *Journal of Chemical Technology Biotechnology*. 1997;70(2):117–140.
51. Mamba G, Mamo MA, Mbianda XY, et al. Nd,N,S–TiO<sub>2</sub> Decorated on Reduced Graphene Oxide for a Visible Light Active Photocatalyst for Dye Degradation: Comparison to Its MWCNT/Nd,N,S–TiO<sub>2</sub> Analogue. *Industrial & Engineering Chemistry Research*. 2014;53(37):14329–14338.
52. Chen X, Mao SS. Titanium dioxide nanomaterials: Synthesis, properties, modifications, and applications. *Chemical Reviews*. 2007;107(7):2891–2959.
53. Sobczynski A, Dobosz A. Water purification by photocatalysis on semiconductors. *Polish Journal of Environmental Studies*. 2001;10(4):195–205.
54. Tian G, Pan K, Fu H, et al. Enhanced photocatalytic activity of S–doped TiO<sub>2</sub>–ZrO<sub>2</sub> nanoparticles under visible–light irradiation. *Journal of Hazardous Materials*. 2009;166(2–3):939–944.



55. Cheng X, Liu H, Chen Q, et al. Construction of N, S codoped TiO<sub>2</sub> NCs decorated TiO<sub>2</sub> nano-tube array photoelectrode and its enhanced visible light photocatalytic mechanism. *Electrochimica Acta*. 2013;103:134–142.
56. Sangareswari M, Sundaram MM. Development of efficiency improved polymer-modified TiO<sub>2</sub> for the photocatalytic degradation of an organic dye from wastewater environment. *Applied Water Science*. 2017;7(4):1781–1790.
57. Liao DL, Badour CA, Liao BQ. Preparation of nanosized TiO<sub>2</sub>/ZnO composite catalyst and its photocatalytic activity for degradation of methyl orange. *Journal of Photochemistry and Photobiology A: Chemistry*. 2008;194(1):11–19.
58. Colmenares JC, Aramendia MA, Marinás A, et al. Synthesis, characterization and photocatalytic activity of different metal-doped titania systems. *Applied Catalysis A: General*. 2006;306:120–127.
59. Xiao Q, Zhang J, Xiao C, et al. Photocatalytic decolorization of Methylene blue over Zn<sub>1-x</sub>CoxO under visible light irradiation. *Materials Science and Engineering: B*. 2007;142(2–3):121–125.
60. Lachheb H, Puzenat E, Houas A, et al. Photocatalytic degradation of various types of dyes (alizarin S, crecein Orange G, methyl red, congo red, methylene blue) in water by UV-irradiated titania. *Applied Catalysis B: Environmental*. 2002;39(1):75–90.
61. Kumar A, Kumar D, Pandey G. Characterisation of hydrothermally synthesised CuO nanoparticles at different pH. *Journal of Technological Advances and Scientific Research*. 2016;2(4):166–169.
62. Fujishima A, Rao TN, Tryk DA. Titanium dioxide photocatalysis. *Journal of Photochemistry and Photobiology C: Photochemistry Reviews*. 2000;1(1):1–21.
63. Matthews RW, McEvoy SR. A comparison of 254 nm and 350 nm excitation of TiO<sub>2</sub> in simple photocatalytic reactors. *Journal of Photochemistry and Photobiology A: Chemistry*. 1992;66(3):355–366.
64. Slokar YM, Marechal AML. Methods of decoloration of textile wastewaters. *Dyes & Pigments*. 1998;37(4):335–356.
65. Baran W, Makowski A, Wardas W. The effect of UV radiation absorption of cationic and anionic dye solutions on their photocatalytic degradation in the presence of TiO<sub>2</sub>. *Dyes and Pigments*. 2008;76(1):226–230.
66. Wang KH, Hsieh YH, Wu CH, et al. The pH and anion effects on the heterogeneous photocatalytic degradation of o-methylbenzoic acid in TiO<sub>2</sub> aqueous suspension. *Chemosphere*. 2000;40(4):389–394.
67. Said Alahiane, Samir Qourzal, Mahmoud El Ouardi, et al. Factors Influencing the Photocatalytic Degradation of Reactive Yellow 145 by TiO<sub>2</sub>-Coated Non-Woven Fibers. *American Journal of Analytical Chemistry*. 2014;5(8):445–454.
68. Gaya UI, Abdullah AH. Heterogeneous photocatalytic degradation of organic contaminants over titanium dioxide: a review of fundamentals, progress and problems. *Journal of Photochemistry and Photobiology C: Photochemistry Reviews*. 2008;9(1):1–12.
69. Asahi R, Morikawa T, Ohwaki T, et al. Visible-light photocatalysis in nitrogen-doped titanium dioxides. *Science*. 2001;293(5528):269–271.
70. Sanghamitra Kundua, Anjali Pala, Anil K Dikshit. UV induced degradation of herbicide 2,4-D: kinetics, mechanism and effect of various conditions on the degradation. *Separation and Purification Technology*. 2005;44(2):121–129.
71. Augugliaro V, Litter M, Palmisano L, et al. The combination of heterogeneous photocatalysis with chemical and physical operations: a tool for improving the photoprocess performance. *Journal of Photochemistry and Photobiology C: Photochemistry Reviews*. 2006;7(4):127–144.
72. Pal B, Kaur R, Grover IS. Superior adsorption and photodegradation of Eriochrome black-T dye by Fe<sup>3+</sup> and Pt<sup>4+</sup> impregnated TiO<sub>2</sub> nanostructures of different shapes. *Journal of Industrial and Engineering Chemistry*. 2016;33:178–184.
73. Liu S, Jaffrezic N, Guillard C. Size effects in liquid-phase photo-oxidation of phenol using nanometer-sized TiO<sub>2</sub> catalysts. *Applied Surface Science*. 2008;255(5):2704–2709.
74. Tang C, Chen V. The photocatalytic degradation of reactive black 5 using TiO<sub>2</sub>/UV in an annular photoreactor. *Water Research*. 2004;38(11):2775–2781.
75. Rajeshwar K, Osugi ME, Chanmanee W, et al. Heterogeneous photocatalytic treatment of organic dyes in air and aqueous media. *Journal of Photochemistry and Photobiology C: Photochemistry Reviews*. 2008;9(4):171–192.
76. Choi H, Stathatos E, Dionysiou DD. Photocatalytic TiO<sub>2</sub> films and membranes for the development of efficient wastewater treatment and reuse systems. *Desalination*. 2007;202(1–3):199–206.
77. Zaleska A. Doped-TiO<sub>2</sub>: A Review. *Recent Patents on Engineering*. 2008;2(3):157–164.
78. Yu JC, Yu J, Ho W, et al. Effects of F- doping on the photocatalytic activity and microstructures of nanocrystalline TiO<sub>2</sub> powders. *Chem Mater*. 2002;14(9):3808–3816.
79. Hong X, Wang X, Cai W, et al. Visible-light-activated nanoparticles photocatalyst of iodine-doped titanium dioxide. *Chem Mater*. 2005;17(6):1548–1552.
80. Chong MN, Jin B, Chow CWK, et al. Recent developments in photocatalytic water treatment technology: A review. *Water research*. 2010;44(10):2997–3027.
81. Kansal SK, Singh M, Sud D. Studies on photodegradation of two commercial dyes in aqueous phase using different photocatalyst. *J Hazard Mater*. 2007;141(3):581–590.
82. Selvam K, Muruganandham M, Muthuvel I, et al. The influence of inorganic oxidants and metal ions on semiconductor sensitized photodegradation of 4-fluorophenol. *Chemical Engineering Journal*. 2007;128(1):51–57.
83. Baran W, Makowski A, Wardas W. The influence of FeCl<sub>3</sub> on the photocatalytic degradation of dissolved azo dyes in aqueous TiO<sub>2</sub> suspensions. *Chemosphere*. 2003;53(1):87–95.