

Theoretical calculation of uv-vis, ir spectra and reactivity properties of tamoxifen drug: a methodology comparison

Abstract

Some theoretical properties were analyzed and compared with the experimental data for Tamoxifen molecule, a drug commonly used as complementary therapy for breast cancer. The molecular structure and some chemical reactivity parameters were calculated through Density Functional Theory using different functionals, including B3LYP, PBE0, PBEPBE, TPSS, TPSSh and the M05 and M06 density functionals of Minnesota family, all of them combined with a 6-31G (d) basis set. The theoretical IR and ultraviolet-visible (V-Vis) spectra were compared with experimental data. Reactivity parameters are of great importance in the pharmacological testing. For this reason, a correlation between the different chemical models was carried out using the electron affinity and ionization potential calculations. According to the outcomes, the methodology that had a better correlation with the experimental data is M06/6-31G (d).

Keywords: DFT, tamoxifen, vibrational analysis

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Abbreviations: TAM, tamoxifen; DFT, density functional theory; SERM, selective estrogen receptor modulator; TD-DFT, time-dependent density functional theory; HOMO, highest occupied molecular orbital; LUMO, lowest unoccupied molecular orbital; EA, electron affinity; I, ionization potential; η , chemical hardness; χ , electronegativity; ω , electrophilicity index; μ , chemical potential

Introduction

Tamoxifen[(Z)-2-(4-(1,2-Diphenyl-1-butenyl)-phenoxy)-N,N-dimethylethanamine] (TAM) has been prescribed to treat patients with breast cancer.¹ TAM is a selective estrogen receptor modulator (SERM).^{2,3} This drug acts as anti-estrogen in breast tissue by interfering with the activity of estrogen: the female sex hormone that promotes the growth of cancer cells in the breast.⁴⁻⁶

On the other hand, it is known that theoretical studies of geometrical, electronic, chemical and spectral properties of drugs are very important as they provide information if the drug has a change in orientation (conformational),⁷ its reactive behavior,⁸ the structure-activity relationship^{9,10} and determine the way the molecule interacts with other species.¹¹

Some authors have studied the structure-activity relationship (QSAR) of TAM,¹² which showed that a small difference in the chemical structure can alter the pharmacological activity.¹³ In recent years, the density functional theory (DFT) calculations reported to provide excellent vibrational frequencies of organic compounds,¹⁴ in spectroscopic properties¹⁵ and reactivity descriptors of organic and inorganic compounds.¹⁶ Also, theoretical studies based on quantum calculations through BLYP/6-31G+(d,p) and B3LYP/6-31G+(d,p) have achieved optimized structures in good agreement with X-ray experimental geometries, and were thus used to investigate the flip-flop process of the TAM propeller and derivatives.¹⁷ However, it is important to find

a suitable methodology that describes the vibrational properties of TAM, since this will lead to a reliable optimized molecular structure. Besides the detailed knowledge about a structure of drugs, spectra and other properties is necessary for better understanding their chemical and biological.¹⁸

The calculation of the vibrational frequencies is a very good aid for the assignment of the IR spectra.¹⁹ UV-Vis spectrophotometry is one of the most frequently employed techniques in pharmaceutical analysis. It involves measuring the amount of ultraviolet or visible radiation absorbed by a substance in solution.²⁰ Furthermore, evaluation of UV-Vis may be necessary in the following circumstances: changes in the synthesis of the drug substance; changes in the composition of the finished product and changes in the analytical procedure.²⁰ The wavelength of maximum absorption (λ_{\max}) is normally selected.

Moreover, to know the absorption bands of the UV-vis spectrum is very important considering because the quantification of the medicinal substance can be measured by λ_{\max} .²⁰ This spectrum also shows the frontier orbital transitions, variations that may lead to the formation of bonds between different atoms. These data are of significant interest for future investigations on the use of carrier vehicles loaded with the drug Tamoxifen or its derivatives.

The main goal of this research was to validate the chemical model that best describes the electronic properties and chemical reactivity parameters of the TAM molecule. Hence, several density functional were selected to calculate the electronic properties, which were compared with experimental results. Thereby, eleven density functionals were evaluated using the 6-31G (d) basis set, reported by Pople et al.²¹ Some of the defined methodologies were chosen by the experience of previous studies of our research group such as: PBE0 in megalol drug²² and chemical reactivity of pyrrole derivatives;²³ B3LYP, M05 and M06 density functionals with Cyanidin molecules²⁴ and

the theoretical spectroscopic analysis of Riluzole, using FT-IR, FT-Raman and UV-Vis techniques were obtained with three functionals (B3LYP, MPWLYP and M06-2X) and 6-311++G(d,p) basis set having agreement between experimental and the calculated.²⁵ According of the study reported by Soto-Rojo et al.²⁶ the M06 functional presents the best alternative to obtain the UV-Vis spectra with good accuracy.²⁶ The theoretical determination of chemical model for TAM can be very useful in the design and development of new and better drugs, allowing so that other researchers have the facility to use the chemical model in other derivatives of the drug tamoxifen without validate the chemical model.

Computational details

All calculations were performed using the Gaussian 09 software.²⁷ The Density Functional Theory (DFT) was used,^{28–30} which has been proven to be a key method for drug design and specifically in the field of transition metals with medicinal applications, and has been successful in terms of accuracy and reliability.³¹ The used functionals were: Perdew, Burke and Ernzerhof GGA functional, PBE0;³² GGA hybrid functionals, Becke three parameter Lee, Yang, and Parr, B3LYP;³³ and PBEPBE;³⁴ Tao, Perdew, Staroverov, and Scuseria meta-GGA, TPSS;³⁵ Truhlar and coworkers meta-GGA local functional M06-L;²⁶ and hybrid meta-GGA functionals M05-2X,^{36,37} M05,^{38,39} M06,^{36,40} M06-2X,³⁶ M06-HF,⁴¹ and TPSSh,³⁵ functionals that have different percentage of Hartree-Fock. These functionals were combined with Pople basis set 6-31G (d)²¹ for the geometrical optimizations, followed by a frequency calculation to confirm that the structure is at the minimum energy state. The theoretical Infrared spectra (IR) were also plotted with the results obtained by the different chemical model frequency calculations and compared with the experimental IR spectrum.

Ultraviolet-visible spectra (UV-Vis) were calculated by solving the time-dependent density functional theory (TD-DFT) equations^{42–44} approach and analyzed through the SWizard program;⁴⁵ theoretical UV-Vis spectra were obtained with different chemical model in the presence of methanol as solvent, using the conductor-like polarizable continuum model (CPCM)⁴⁶ which defines the cavities as envelopes of spheres centered on atoms or atomic groups. Inside the cavity the dielectric constant is the same as in vacuum, outside it takes the value of the desired solvent;⁴⁷ the equations were solved for 20 excited states.

The conceptual aspect of the density functional method has been amply used to understand the chemical reactivity, to define a set of chemical concepts such as the electron affinity (EA),^{48,49} ionization potential (I),^{48,49} chemical hardness (η),⁵⁰ electronegativity (χ),^{51,52} electrophilicity index (ω)⁵³ and chemical potential (μ),⁵³ which are linked to their electronic structure and chemical reactivity. There are different ways to calculate the reactivity parameters. One formulations consists in calculate the ionization potential and electron affinity as the negative of the HOMO and LUMO energy respectively. It is called Koopmans' theorem. This approximation is widely used with good results.^{54,55} In the research group of the authors is commonly to calculate with the second approximation, it uses the energy difference between the precursor molecule and the species formed by the adding or removing of an electron. It is the energy difference approximation and considers the energy (E) as a function of the number of electrons (N). The equations of the chemical reactivity calculated by this method are shown in Table 1.

Table 1 Equations of global reactivity

$EA = E(N) - E(N+1)$	48,49
$I = E(N-1) - E(N)$	48,49
$\eta = \frac{(I-EA)}{2}$	50
$\chi = \frac{(I+EA)}{2}$	51,52
$\omega = \frac{\mu^2}{2\eta}$	53
$\mu = -\chi$	53

Results and discussion

Geometry optimization and Calculated IR spectra

The TAM structure was optimized in gas phase with all density functionals mentioned above, followed by a frequency calculation to confirm that the molecule has a minimum energy conformation, allowing to obtain the theoretical IR spectra, which were compared with experimental results,⁵⁶ to choose a reliable functional. Figure 1 shows the optimized TAM structure with M06/6-31G (d) methodology.

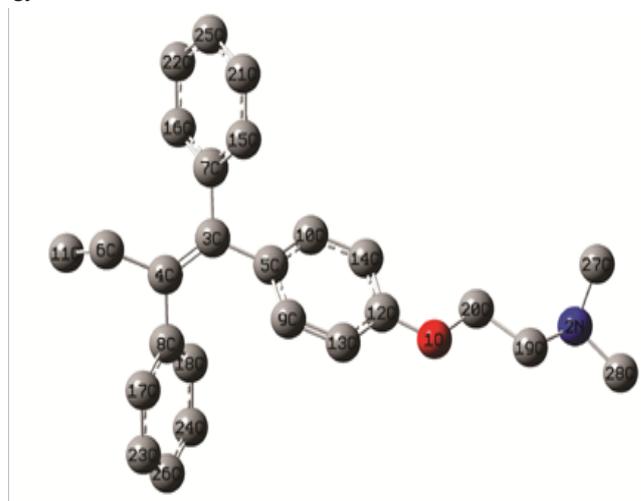


Figure 1 Optimized geometry of Tamoxifen obtained with M06/6-31G (d).

The optimized structure of TAM shows a non-planar geometry because it has three dihedral angles: C5-C3-C7-C16 with a value of 126.40 degrees, the C5-C4-C3-C8 with -9.32 degrees, and C3-C4-C5-C6 with 171.69 degrees, and a dihedral angle at the opposite end of the molecule, which corresponds to the angle C20-C19-N2-C27 with a value of 69.86 degrees.

The theoretical IR spectra values obtained with the different chemical models were compared with the experimental FT-IR spectrum values. Some of the principal vibrations correspond to different func-

tional groups, such as: Alkane C-H stretching, Alkene -C=C- stretching and C=C ring stretching. The frequency values found for each functional are shown in Table 2. These data were analyzed and the chemical models that reproduce with the best accuracy the experimental data reported by Shivam,⁵⁶ are M06/6-31G (d) and PBE0/6-31G (d). Figure 2 & 3 show the theoretical Tamoxifen IR spectra for M06/6-31G (d) and PBE0/6-31G (d) respectively.

UV-Visible absorption spectra

The maximum UV-Vis absorption spectra results for Tamoxifen were performed in the presence of solvent, under the same optimiza-

tion conditions mentioned above. The calculated maximum absorption wavelengths (λ_{\max}), the oscillator strength (f), and electronic transitions are shown in Table 3. According to the results obtained in the UV-Vis theoretical spectra the chemicals models that exhibit a better agreement, are M06/6-31G (d) with 27% HF exchange underestimate the λ_{\max} by 6 nm, PBE0/6-31G(d) with 25% HF exchange underestimate by 7.8 nm and M05/631G(d) with 25% HF exchange underestimate with 1.8 nm in relation to the 298.2 nm maximum absorption wavelength experimental reported by Merey.⁵⁷ Figure 4 shows the theoretical Tamoxifen UV-Vis spectra for M06, PBE0 and M05 all 6-31G (d) basis set.

Table 2 Comparison of the theoretical vibrations modes with experimental FT-IR spectrum of Tamoxifen

Chemical model	Vibrations		
	Alkane C-H Stretching cm^{-1}	Alkene -C=C- Stretching cm^{-1}	C=C ring stretching
B3LYP/6-31G(d)	3029.13	1655.32	1542.14
M05/6-31G(d)	3008.68	1680.05	1480.29
M05-2X/6-31G(d)	3014.92	1726.33	1515.79
M06/6-31G(d)	2929.74	1690.54	1458.97
M06-2X/6-31G(d)	3061.26	1722.71	1502.91
M06-L/6-31G(d)	2950.85	1655.47	1399.23
M06-HF/6-31G(d)	3032.25	1747.4	1519.65
PBE0/6-31G(d)	2967.06	1687.72	1496.96
PBEPBE/6-31G(d)	2957.05	1596.74	1513.66
TPSS/6-31G(d)	3037.74	1607.86	1515.4
TPSSh/6-31G(d)	3027.43	1640.92	1539.88
Experimental data ⁴⁸	2800-3000	1700-1740	1400-1500

Table 3 Comparison of the wavelengths (λ_{\max}), modes with experimental UV-Vis spectra of Tamoxifen calculated with different chemical model with the 6-31(G) basis set

Chemical model	λ_{\max}	F	Electronic transitions
B3LYP	310.58	0.4047	H-0->L+0(+98%)
M05	300.01	0.4507	H-0->L+0(+98%)
M05-2X	274.3	0.578	H-0->L+0(+96%)
M06	304.54	0.4536	H-0->L+0(+98%)
M06-2X	277.63	0.56	H-0->L+0(+96%)
M06-L	358.4	0.2268	H-0->L+0(+75%) H-1->L+0(21%)
M06-HF	246.08	0.6775	H-0->L+0(+87%)
PBE0	306.07	0.448	H-0->L+0(+99%)
PBEPBE	410.22	0.0124	H-1->L+0(+74%) H-0->L+0(26%)
TPSS	379.75	0.0295	H-1->L+0(+80%) H-0->L+0(20%)
TPSSh	333.76	0.3522	H-0->L+0(+97%)
Experimental data ⁵⁷	298.2		

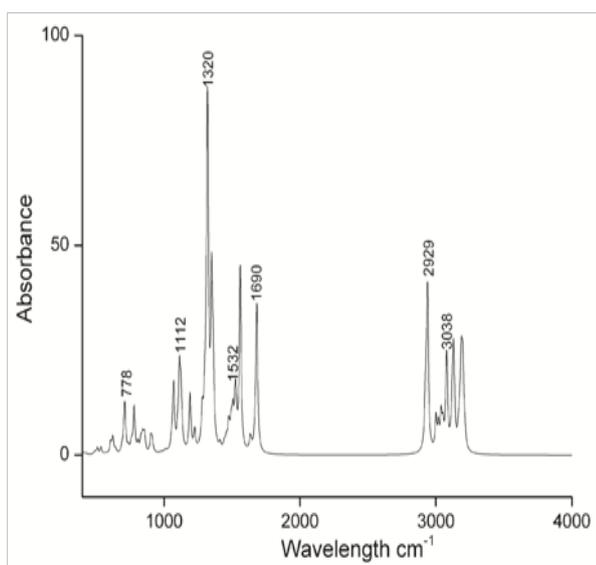


Figure 2 Theoretical IR spectra of Tamoxifen at the M06/6-31G (d) level of calculation.

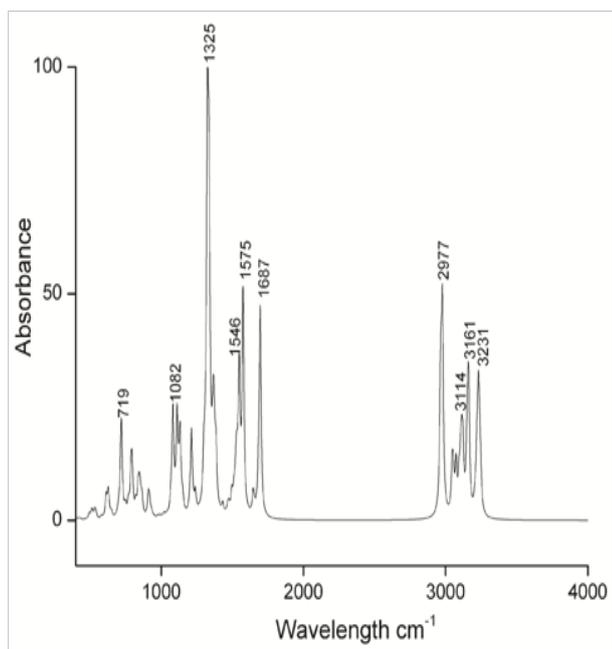


Figure 3 Theoretical IR spectra of Tamoxifen at the PBE0/6-31G (d) level of calculation.

Variables design and statistical analysis

The chemical models M06/6-31G(d) and PBE/6-31G(d) reproduce adequately the experimental FT-IR spectrum and there is a good approach to the experimental maximum absorption wavelength with M06/6-31G(d), PBE/6-31G(d) and M05/6-31G(d). Additionally, some reactivity parameters calculation were developed with the eleven functionals mentioned in the computational details section, with the aim to obtain variables that led to develop a statistical analysis of the methodologies and define the best one.

The developed calculated values are electron affinity (EA) and ionization potential (I) and the results are reported in Table 4. In this

case all the electron affinity values are negative which means a resonance state or an accepted unbound electron according to Lewars, who states that the electron affinity has negative values when the accepted electron is ejected in microseconds.⁴⁹

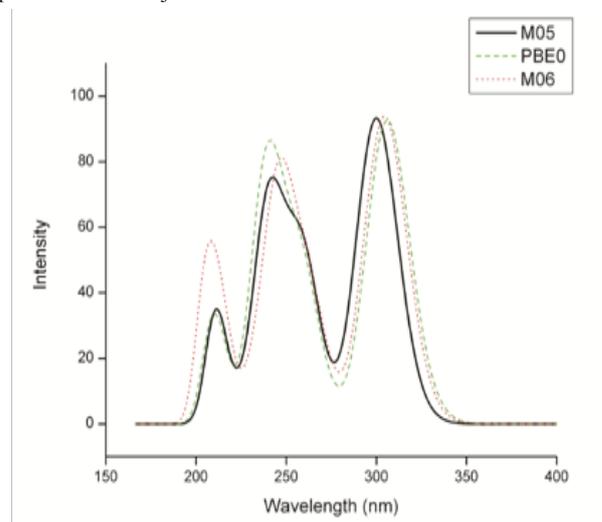


Figure 4 UV-Vis spectra of Tamoxifen calculated with M06/6-31G (d), PBE0/6-31G (d) and M05/6-31G (d).

Table 4 Ionization Potential and Electron Affinity of Tamoxifen with different functionals and 6-31G (d) basis set

Functional	EA (eV)	I (eV)
B3LYP	-0.62	6.55
M05	-0.69	6.6
M05-2X	-0.73	7.03
M06	-0.56	6.67
M06-2X	-0.67	7.05
M06-L	-0.4	18.8
M06-HF	-0.91	7.56
PBE0	-0.56	6.65
PBEPBE	-0.27	6.15
TPSS	-0.38	6.15
TPSSh	-0.48	6.36

Regarding the ionization potential the values are in a range of 6.15 to 7.56eV with the exception of the value calculated with M06L, with a value of 18.8eV. This can be attributed to the fact that M06L is the only local functional, with 0% HF exchange, and considering that HF underestimates the ionization potential.⁵⁸ The only reactivity value found in Tamoxifen is the ionization potential, calculated with B3LYP by Garrido with a value of 6.49eV.⁵⁹ In this case, the results obtained with B3LYP, M05, M06 and PBE0 are the closest to the value calculated by Garrido.

A dispersion analysis was done with the results of the reactivity values obtained. It was eliminated the Ionization Potential calculated with M06-L/6-31G(d), due to the different value compared with the others.

The mean value for the scattering data obtained is -0.59eV and

6.70eV for EA and I respectively, while the median has values of -0.59eV for EA and 6.63eV for I. These results allowed to calculate the standard deviation for the EA, which exhibits a value of 0.17 and 0.41 for I, and coefficients of variation of 28% and 6.3% respectively. These values represent excellent quality and low variability,⁶⁰ and contribute to define the best model chemistry. The dispersion analysis results are in Figure 5 where the dash blue line represents the mean value for both, EA and I.

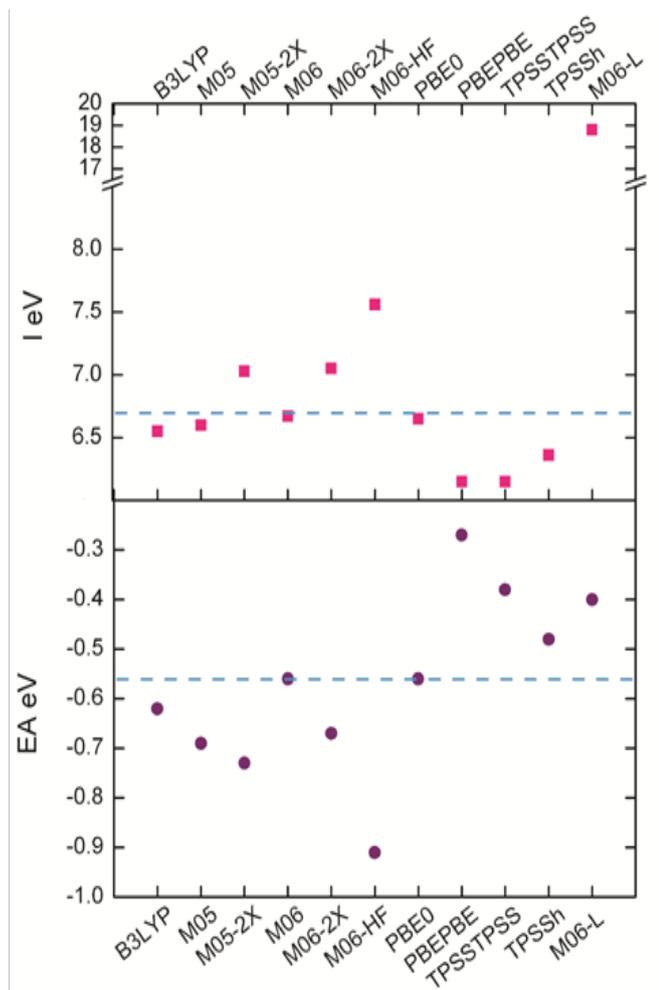


Figure 5 Ionization potential and electron affinity for the model chemistries used in this study and the dispersion between the values.

The dispersion results show that there are two chemistry models with a first-rate approximation for being considered as the best methodology, M06/6-31G (d) and PBE0/6-31G (d). Considering that these two methodologies also have the best agreement with the experimental vibrations and the maximum absorption wavelengths shown above, a third statistical function, the z-score, was used to define the best methodology. This function indicates how many standard deviations an element is from the arithmetic mean.⁶¹

According to the results for z-score which are in Table 5, the typical value of the electron affinity scores (zAE) for M06/6-31G (d) and PBE0/6-31G (d) is the same, 0.18. This indicates that both functionals have the same relative position to the mean value. However, typical ionization potential scores (zI) for the chemical models M06/6-31G (d), and PBE0/6-31G (d) are different, being -0.07 and -0.12 respectively.

Based on these results M06/6-31G (d) is the best approximation to calculate the Tamoxifen properties. We propose that the chemical model M06/6-31G (d) presents excellent approaches in this analysis.

Table 5 Z score for ionization potential and electron affinity of M06/6-31G (d) and PBE0/6-31G (d) methodologies

Functional	AE (eV)	I (eV)	zEA	zI
M06	-0.56	6.67	0.18	-0.07
PBE0	-0.56	6.65	0.18	-0.12

Theoretical characterization of tamoxifen with M06/6-31G (d)

Once the methodology was defined, it was developed, the characterization of the molecule of Tamoxifen that included the structural parameters of minimal energy geometry and a comparison versus experimental bond lengths and angles. Also the electronic properties as electrostatic potential and reactivity parameters.

Structural parameters: We have compared the chemical model M06/6-31G (d) optimized geometry of TAM to the X-ray data⁶² (Table 6). The analysis the difference between the experimental and theoretical of bond lengths shows that in C-O bond oscillate 0.054Å, whereas N-C the difference is 0.063Å and in C-C is 0.063Å. On the other hand in the bond angles the difference in the C-O-C has 7.53°, in C-N-C 1.29°, while alkene functional group with 0.512°, 0.077° and -0.659° in the bonds angles C3-C5-C9, C3-C7-C15 and C4-C8-C18 respectively. Therefore we have pointed out the good agreement of our theoretical results and the experimental ones.

Electrostatic potential surface: To evaluate the electronic distribution around Tamoxifen, the electrostatic potential surface (EPS) was plotted (Figure 6). The map show the negative potential sites that include electronegative atoms as well as the positive potential sites located around the hydrogen atoms. The main negative center includes the oxygen atom of the ether group and the nitrogen of the amine group. These atoms could be responsible of the interaction through the hydrogen bond with the active site in hormone receptors. This is because this type of bond involves a hydrogen atom that is attracted to a strongly electronegative atom, such as oxygen, fluorine, or nitrogen of a polar covalent bond, in the same or another molecule.⁶³ The contour around TAM is found in the functional groups aryl, ethyl and ether of the drug; these regions most probably constitute the active site of the drug in chemical reactions. The EPS analysis agrees with the frontier molecular orbitals study, where highest occupied molecular orbital (HOMO) and the lowest unoccupied molecular orbital (LUMO) are located over the phenyl, ethyl and ether groups (Figure 7).

Chemical reactivity: The chemical reactivity parameters obtained for TAM drug are: ionization potential (I), which is defined as the energy needed to remove an electron from a molecule by 6.67eV; electron affinity (EA), which measures the ability of a molecule to accept electrons or form anions upon -0.56eV; chemical hardness (η), is the resistance to change its electronic configuration with 3.61eV; electronegativity (χ), represents the tendency of atoms or molecules to attract electrons with 3.05eV; electrophilicity index (ω), that gives an idea of the stabilization in energy when the system acquires electrons from the environment up to saturation by 1.29eV; and chemical potential (μ), measures the tendency of electrons to be released from a system is -3.05eV. Which they were calculated using the molecular energies of the cationic, ionic and neutral species (Table 7).

Table 6 The optimized geometrical parameters (Bond lengths (Å) and angles (°) and experimental of Tamoxifen

Bond Lengths (Å)				Bond Angles (°)				
		X-Ray	M06/6-31G(d)			X-Ray	M06/6-31G(d)	
1O	12C	1.4	1.357	12C	1O	20C	125.7	118.165
1O	20C	1.477	1.411	19C	2N	27C	114	112.031
2N	19C	1.547	1.45	19C	2N	28C	112.6	110.377
2N	27C	1.485	1.449	27C	2N	C28	109.7	109.996
2N	28C	1.505	1.448	4C	3C	5C	127.4	123.047
3C	4C	1.5	1.355	4C	3C	7C	121.8	122.493
3C	5C	1.605	1.485	5C	3C	7C	108.9	114.436
3C	7C	1.525	1.487	3C	4C	6C	127.4	123.667
4C	6C	1.569	1.508	3C	4C	8C	120.3	122.081
4C	8C	1.549	1.486	6C	4C	8C	112.3	114.245
5C	9C	1.443	1.403	3C	5C	9C	122.6	122.088
5C	10C	1.457	1.396	3C	5C	10C	119.7	120.368
6C	11C	1.535	1.526	9C	5C	10C	117.7	117.506
7C	15C	1.452	1.4	4C	6C	11C	115.2	113.054
7C	16C	1.453	1.399	3C	7C	15C	120.8	120.023
8C	17C	1.436	1.399	3C	7C	16C	121	121.75
8C	18C	1.427	1.4	15C	7C	16C	118.2	118.184
9C	13C	1.435	1.382	4C	8C	17C	120.5	121.159
10C	14C	1.406	1.391	4C	8C	18C	120.1	120.602
12C	13C	1.454	1.398	17C	8C	18C	119.4	118.172
12C	14C	1.46	1.394	5C	9C	13C	121.7	121.323
15C	21C	1.398	1.388	5C	10C	14C	121.9	121.947
16C	22C	1.427	1.391	1O	12C	13C	115.7	115.961
17C	23C	1.414	1.39	1O	12C	14C	125.3	124.518
18C	24C	1.414	1.389	13C	12C	14C	119	119.52
20C	19C	1.529	1.514	9C	13C	12C	119.5	120.248
21C	25C	1.43	1.392	10C	14C	12C	120.2	119.441
22C	25C	1.423	1.391	7C	15C	21C	121.3	120.954
23C	26C	1.419	1.391	7C	16C	22C	119.9	120.967
24C	26C	1.425	1.392	23C	17C	8C	120.5	120.952
				8C	18C	24C	120	120.956
				2N	19C	20C	111.1	112.469
				1O	20C	19C	111.6	106.643
				15C	21C	25C	120	120.193
				16C	22C	25C	120.2	120.145
				17C	23C	26C	119.6	120.187
				18C	24C	26C	120.2	120.229
				21C	25C	22C	120.4	119.54
				23C	26C	24C	120.4	119.491

Table 7 Chemical reactivity parameters with M06/6-31G (d)

Drug	Chemical reactivity descriptors					
	I (eV)	AE (eV)	χ (eV)	η (eV)	μ (eV)	ω (eV)
TAM	6.67	-0.56	3.05	3.61	-3.05	-1.29

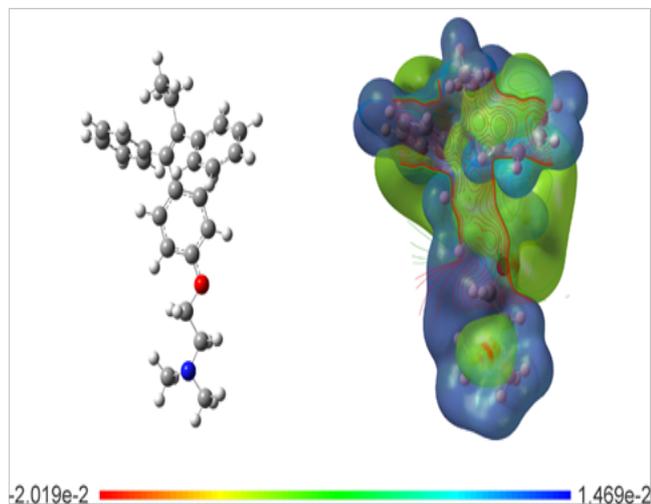


Figure 6 Optimized structure and calculated electrostatic potential map on the molecular surface for TAM drug calculated at the M06/6-31G (d) level of theory. Color range oscillates -2.019×10^{-2} to 1.469×10^{-2} ; blue, more positive; red more negative.

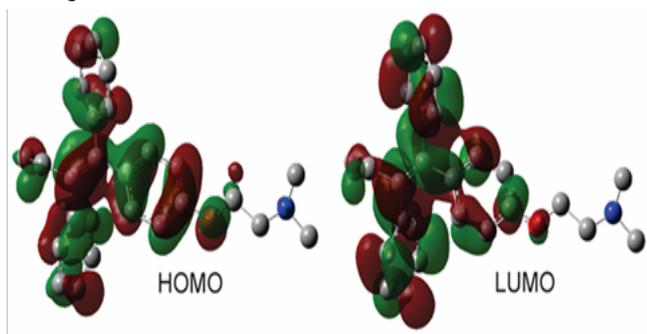


Figure 7 Molecular orbitals for TAM drug calculated at the M06/6-31G (d) level of theory.

Conclusion

In this work, different chemical models were implemented to determine the structural parameters and reactivity of TAM. The M06/6-31G (d) and PBE0/6-31G (d) chemical models had the best approach to the FT-IR spectrum for maximum absorption wavelengths the best results were obtained with M06/6-31G(d), PBE0/6-31G(d) and M05/6-31G(d). These chemical models yielded better results for modeling the structural parameters studied. The performed statistical analysis of the electron affinity and ionization potential reactivity parameters showed that has the closest approximation to the mean value is for M06/6-31G(d). Therefore, it was defined as the best methodology. The calculation of the structural parameters with this methodology was in good agreement with the experimental values and the electrostatic potential surface determined which moieties of the molecule will interact with the active site of the hormone receptor.

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Conflict of interest

The author declares no conflict of interest.

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