

Investigating regional groundwater flow systems in baja california central desert region

Abstract

In Northern Baja California, environmental conservation may not be fully accomplished in the absence of a clear understanding of key factors controlling and protecting biodiversity. Groundwater is an environmental agent triggering the presence of a wide range of arid ecosystems; therefore, any sustainable development in the region has to maintain a balance between environmental, social and economical benefits. The present study has placed a great emphasis on the understanding and distribution of groundwater flow systems, and their environmental interactions in the protected area of the Valle de Los Círios, using physical-chemical and isotope methods. Results show the presence of three groundwater groups portraying contrasting flow systems conditions. Analyzed samples have been classified as fresh-brackish, and thermal alkaline groundwater as suggested by pH, temperature and TDS. The Br/Cl ratio in most samples was close to the ocean molar value (1.5×10^{-3}) indicating the importance of marine aerosol fraction in the recharged waters; however, Mg/Cl, Na/Cl, Ca/Cl, SO_4/Cl and $\text{Cl}/\text{HCO}_3 + \text{Cl}$ ratios were significantly different to those of seawater mixing conditions. Dominant ions (Na, Ca and Cl), chemical age proxies (Li), and $\delta^{18}\text{O}$ suggest waters with long residence time, with the possibility of been recharge during the late glacial period. Thermal ground water shows low minor and trace element content and exhibits the dominance of Na-plagioclase hydrolysis; their low levels of Sr and Ba suggest negligible contact with calcareous rocks and their estimated temperature at depth (80 to 100°C) and TDS suggest groundwater has travelled through fractured media. This introduction study suggests that desertic species in Valle de Los Círios are highly sustained by groundwater of local and intermediate flows.

Keywords: Geochemical analyses, Stable Isotopes, Flow systems, Valle los Círios, Northern Baja California, Mexico

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Introduction

The Baja California peninsula has long been recognized as a hotspot for plant richness and endemism and has been the subject of 14 decrees as protected areas.¹ However, its extraordinary and unique diversity is not adequately protected by the existing protected areas protocols.¹ In the absence of perennial surface-water sources, any economic development in the area will depend solely on groundwater. It is therefore substantial to establish a new approach that broadens the goal of biodiversity protection. Various studies²⁻⁵ have shown the groundwater role in controlling ecological functioning and scatter vegetation, especially in its discharge zones. Thus, hydrological processes are the main controlling factor in the ecological function and variation of uncultivated lowlands.² These processes are related to the evolving hierarchy of groundwater flow systems. Since the encounter of the flow systems concept,⁶ the hydrogeology's basic paradigm has shifted from confined flow in aquifer units to cross-formational-flow in drainage basins.⁷ Consequently, groundwater has been recognised as a fundamental geologic agent, generating and modifying natural processes and phenomena of scientific, practical and economic interest.⁷ The distribution of vegetation cover in response to nutrient and moisture conditions is generated by groundwater in-flow and out-flow; water level fluctuation; negative water-balance in recharge zones, and water surplus in discharge zones (wetland); anionic changes with depth and along groundwater flow systems; negative and positive geothermal anomalies, are considered to be the main indicators of groundwater flow hierarchy, among others.⁷ The

use of physical-chemical hydrogeological approaches in parallel with a wide range of geological settings and field parameters remain a cornerstone in understanding, delineating and defining groundwater flow patterns and in establishing regional flow models. Several studies⁸⁻¹¹ have shown the positive correlation between groundwater residence time and the concentration of certain major and trace elements as Li. Nevertheless, groundwater age determination remains a challenge when it comes to: the use of plug-flow and binary mixing models, the interpretation of environmental tracers in fractured rocks, the use of radiocarbon in environment dominated by volcanic rocks, and the age sensitivity to vertical heterogeneity.¹² For this purpose, it is fundamental to understand the behaviour of chemical and environmental tracers in their geological, morphological, and climate contexts. In the Central Desert region of Baja California, studies dedicated to regional understanding of groundwater flow functioning are infrequent; development and population growth is scarce so the availability of groundwater data is limited to scattered springs and few dug wells. Therefore, due to the scarcity of hydraulic and chemical data this work aims to meet the following objectives under the Groundwater Flow System methodology:

- i. To analyse groundwater-rock interaction processes using physical-chemical and isotope methods
- ii. To chemically differentiate groundwater flow patterns and distribution of discharge zones by understanding the theoretical hydrogeological presence of regional, intermediate and local flows within the ecosystem of the Baja California region.

Study region

The study region is located in the northern part of Baja California peninsula, in the Playas de Rosarito District; it extends over a surface of 6,000 km² (Figure 1), between latitude 28° 0' 49.368", 28° 59' 28.752" N and longitude 112° 44' 45.52", 114° 26' 54.096" W. The region is enclosed between the Libertad and San Borja mountains range and

it includes the protected area of Valle Los Cirios with 164 endemic species in a distinct arid environment. The peninsula is covered with 20 different types of climate; from very arid to temperate.¹³ Most of Baja California has a mean annual temperature above 18 °C and a mean annual rainfall lower than 200 mm. The highest rainfall (500–700 mm) occurs in the highlands of Sierras of San Pedro Mártir and La Laguna, in both latitudinal extremes of the peninsula.¹

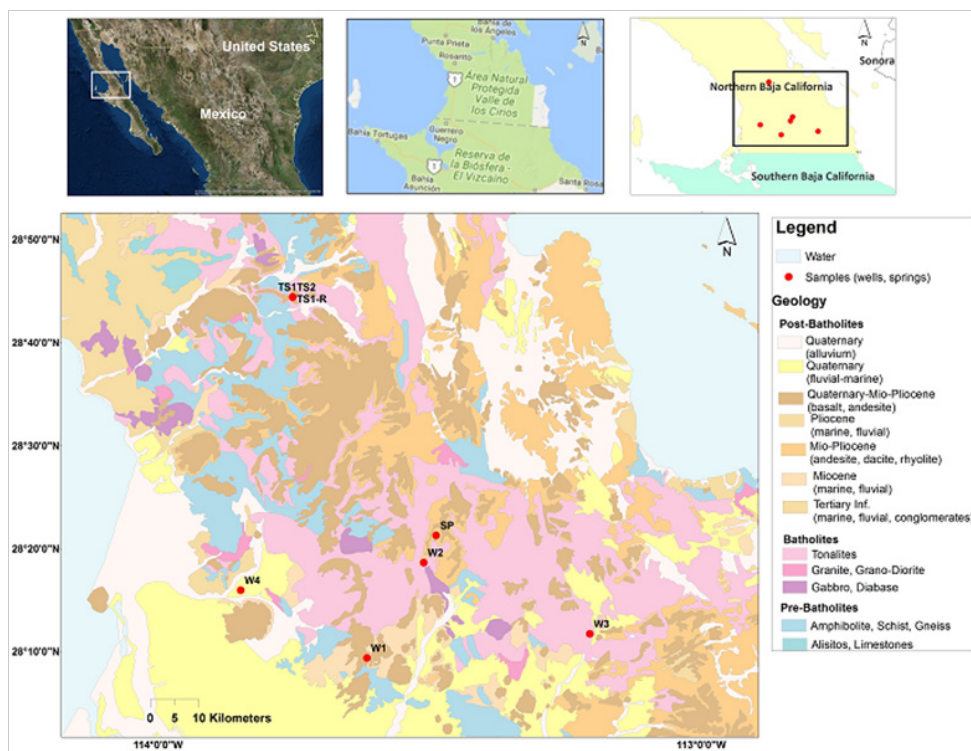


Figure 1 Map showing outcropping geological formations and location of the study region within the Northern Baja California Peninsula.

Geological setting

The geological substratum that supports the rich marine and terrestrial biodiversity and distinctive ecological niches in the Gulf of California region is predominantly composed of granitic and volcanic rocks, as well as sediments derived from these igneous rocks of various ages.¹⁴ The rocks of Baja California (Figure 1) contain the record of two geologic revolutions; the Mid-Mesozoic event which occurred from Jurassic through Middle Cretaceous time and is documented by volcanic strata and the emplacement of pervasive granitic rock; the second, the Mid-Cenozoic event which involved the accumulation of a wide variety of volcanic rocks, a granite emplacement, and the creation of the Gulf of California.¹⁵ These two events divide the geological history of Baja California into significant intervals; what occurred before the culminating granitic emplacements of Middle Cretaceous time is referred to as “pre-batholithic”; everything thereafter, as “post-batholithic”. Along the western shore of Baja California there is a belt consisting of Mesozoic volcanic and volcano-clastic rocks (Figure 1). Cretaceous intrusions in the western Peninsular Ranges Batholith (PRB) were emplaced about ca. 140 Ma.^{16,17} Compositions range from gabbro to rare granite that includes low K quartz gabbro to quartz diorite. The north-western PRB includes gneiss-granite in a belt that extends at 150 km northbound, beyond the international

USA border.¹⁸ In the central part of the peninsula there is a belt of metamorphosed shale and sandstone (Figure 1). On the eastern side there is a belt that contains a great variety of Palaeozoic meta-sedimentary rocks, including those derived from carbonates rocks, metamorphosed shale, sandstone, conglomerate, and limestone, with some volcanic rocks, and weakly metamorphosed pebbly mudstone, wacke, quartzite, thick carbonate units, banded chert, marlstone, and coarse arkose¹⁵ (Figure 1). This eastern part includes mid-Cretaceous intrusions that occurred between ca. 105 and ca. 90.¹⁶ The eastern PRB varies in composition from gabbro to rare granite but are dominated by tonalite-trondhjemite-granodiorite or biotite tonalite.^{15,17–21} These rocks are lower in dark minerals and potash feldspar (low in K₂O, FeO, and MgO, and high in SiO₂ and Al₂O₃) than the rocks of the western part of the batholith.²² Marine Palaeocene and early Eocene strata occur in many places along and south from the Western Cape territory of Baja California.^{23,24} The marine Palaeocene and Eocene strata are deltaic or near-shore deposits. There are also extensive deposits of non-marine, post-batholithic, pre-volcanic sedimentary rocks of undetermined age.¹⁵ The marine Pliocene deposits of southern and northern Baja California are littoral sandstone and conglomerate derived from the basement rock and older sedimentary strata immediately to the east. The Quaternary (Pleistocene-Holocene) deposits are characterized by alluvial, marine and non marine deposits (Figure 1).

Methodology

The methodology adopted in this study highlights the role of groundwater as a geologic agent by applying the gravity driven flow concept^{6,7,25} which differentiates the presence of three major type of flow systems (local, intermediate, regional), based on the physical-chemical and environmental isotopes evolution, and regional natural indicators (discharge zones, type of soil and vegetation). Eight groundwater samples were collected in November 2013 from wells and springs at the lower parts of several canyons in the Libertad mountain range, running deep into the Central Desert region and constituting one of the most interesting transition and controversial zones for vegetation experts.²⁶ The Libertad mountain range is a major transition between two desert regions, the Central Desert towards the north, and the Vizcaino Desert to the south and is included in the protected area of Valle de Los Cirios.²⁶ Major difficulties were encountered during sampling due to the lack of access to operational wells in the area of El Arco, the presence of abandoned wells, and the absence of monitoring wells. Operational wells were found to be sealed and lithology-logs were absent. When water outlet allowed, water sampling was carried out using a flow-cell to avoid contact between groundwater and atmosphere. Spring-waters were collected by directly dipping the sampling bottles and submerging electrodes. Physical-chemical parameters like temperature, electrical conductivity (EC), pH, redox potential (Eh), dissolved oxygen (DO), and alkalinity were measured *in situ*. Equipment and electrodes used were calibrated accurately against known standards before and during related field measurements. Polyethylene bottles of 60 ml were pre-sterilized by 10% HNO₃ immersion and distilled water wash to avoid contamination; during sampling they were rinsed three times with sampled water. Samples were filtered using a 0.45 µm cellulose acetate membrane. The carbonates alkalinity was determined using a digital titrator acid-base titration (H₂SO₄, 1.6N). The HCO₃ concentration was deduced using Gran method. Sampled groundwater was analysed for anions, cations, trace elements, nitrates and stable isotopes (δ²H, δ¹⁸O). Cation samples were acidified with high purity concentrated HNO₃ to reach samples a pH≈2. For nitrate samples, sulphuric acid was added. All samples were preserved at low temperature (≈4°C). Chemical and isotope analyses were conducted in the Actlabs Group, Canada and in the Laboratory of Isotope Geochemistry, University of Arizona, Tucson (USA), respectively. Major and minor cations were analysed using the inductively coupled plasma-mass spectrometry (ICP-MS),

and the anions (F, Cl, NO₃, NO₂, Br, SO₄) analyses were performed using Ion Chromatography system (IC). Geochemical speciation modelling and saturation index calculations were carried out using the PHREEQC code.²⁷

Results and discussions

The interpretation of the hydro-chemical results could be affected by uncertainty on data reliability, to ensure the quality of carried analyses, the ionic balance was calculated and duplicate sample was used. Error values were from 0.08% to 7.81% (Table 1); suggesting good accuracy of the given analyses. Groundwater in the sampled wells and springs could be initially classified from fresh to brackish according to total dissolved solids (TDS) values (307-3,002 mg/L). The salinity of domestic wells (W₁, W₂, W₄), and spring (S_p) (Table 1) was indicated by the importance of TDS and the ionic strength, which ranges between 0.01 and 0.06 M; this implies that cations and anions are more soluble in these waters due to the importance of electrostatic interactions among ions. In general, Na and HCO₃-Cl are the dominant ions, and the groundwater composition can be displayed as follows: Na>Ca-Cl>HCO₃ for W₁, Na>Ca-Cl for W₂ & W₄, Na>Ca-HCO₃ for S_p, Ca-Na-HCO₃ for W₃ & Na-Cl-HCO₃ for TS₁, TS₂ & TS₁-R (=TS₁ duplicate) (Table 1). The relation between lithology and groundwater composition is represented in a Piper diagram (Figure 2), in which the presence of three major water groups may suggest different lithology, residence time and weathering conditions (Grp1: TS₁, TS₂ & TS₁-R; Grp2: W₁, W₂, W₄ & Grp3: W₃ & S_p (S_{sp}: seawater composition in Bahia Concepcion).²⁸ The anion content depicts the importance in the progress of water-rock interaction. Group 2 represents the more evolved waters as compared to Grp3 and Grp1. The Grp1 shares a salt-fresh water mixture, whereas the fresh water in Grp3 had a cation-exchange signature. According to the concentration of TDS and Cl, waters in Grp2 can be classified as water having an intensive geochemical interaction with the material of the aquifer unit. The igneous and metamorphic rock types of the study region are well represented in the groundwater cation chemistry which suggest a chemical weathering of ferromagnesian (basalt) and feldspars rich plagioclase rocks (tonalite, granite, granodiorite). The groundwater in the study region was initially classified in three groups; in which a simple mixture (mixing line) among samples is unlikely to occur due to distinctive physical-chemical (pH, T, TDS, δ²H, δ¹⁸O), geological, and recharge conditions (altitude, climate) characteristics among samples.

Table 1 Physical and chemical analyses of groundwater samples

Sample ID	Sampling Date	Water type	X(m)	Y(m)	Elevation m(asl)	T (°C)	pH	Eh (mV)	DO (mg/l)	EC (µS/cm)	TDS (mg/l)	Ca (mg/l)	Mg (mg/l)	Na (mg/l)	K (mg/l)	Cl (mg/l)	Alkalinity (mg/l)	HCO ₃ (mg/l)	CO ₃ (mg/l)	SO ₄ (mg/l)	NO ₃ (as N) (mg/l)	NO ₂ (as N) (mg/l)	IB*** %
Detection limit											700	2	5	30	0.03					0.03	0.01	0.01	
W1 (D)	26/11/2013	Na>Ca>Mg<Cl>HCO ₃	-12648000	3268596.473	261	23.4	7.16	105	4.4	1547	1037.69	111.0	40.0	169	6.75	388	201.11	244.53		46.1	< 0.05	2.45	1.20
W2 (D)	27/11/2013	Na>Mg>Ca<Cl	-12636000	3288159.875	448	25.6	6.78	116	3.6	4590	3002.28	273.0	188.0	453	12.2	1320	423.44	516.28		195.0	< 0.3	0.70	0.72
SP (Sp)	27/11/2013	Na>Ca>Mg<HCO ₃ >Cl	-12633000	3293738.581	577	20.4	6.96	-40	11.3	1488	997.49	82.2	43.1	144	5.93	173	403.28	491.04		18.0	< 0.04	< 0.04	2.70
W3 (D)	29/11/2013	Ca>Na>Mg<HCO ₃ >Cl	-12602000	3273560.096	505	23.5	7.22	173	2.7	495	310.23	36.7	10.6	33.7	5.91	32.5	128.48	155.79		11.0	< 0.02	4.64	6.66
W4 (D)	29/11/2013	Na>Ca>Mg<Cl	-12673000	3282495.712	81	20.8	7.02	87	1.9	2310	1504.04	147.0	75.1	244	4.86	602	227.87	277.66		126.0	< 0.1	2.02	0.08

Table Continued....

Sample ID	Sampling Date	Water type	X(m)	Y(m)	Elevation m(alt)	T (°C)	pH	Eh (mV)	DO (mg/l)	EC (µS cm)	TDS (mg/l)	Ca (mg/l)	Mg (mg/l)	Na (mg/l)	K (mg/l)	Cl (mg/l)	Alkalinity (mg/l)	HCO ₃ (mg/l)	CO ₃ (mg/l)	SO ₄ (mg/l)	NO ₃ (as N) (mg/l)	NO ₃ (as N) (mg/l)	IB*** %
TS1 (TSp)	30/11/2013	Na-Cl- HCO ₃	-12663000	3342616.458	428	34.4	9.31	-217	3.5	512	304.99	4.0	0.068	101	1.15	78.8	70.19	61.7	9.9	26.6	< 0.02	< 0.02	5.85
TS2 (TSp)	30/11/2013	Na- HCO ₃ -Cl	-12663000	3342570.611	436	27.8	9.50	-3	3.9	512	326.77	4.0	0.048	101	1.18	78.9	90.19	80.5	13.9	27.1	< 0.02	< 0.02	0.61
TS1-R (Rp-TS1)	30/11/2013	Na-Cl- HCO ₃	-12663000	3342616.458	428	34.4	9.31	-217	3.5	512	307.83	4.0	0.044	105	1.14	78.3	70.19	61.7	9.9	26.7	< 0.02	< 0.02	7.81

D, Domestic well; Sp, Spring; TSp, Thermal Spring; Rp, Replicate; *TDS, Calculated with Aquachem; **, CaCO₃; §, Field Measured; EC, Electrical Conductivity; IB***, Error ionic balance

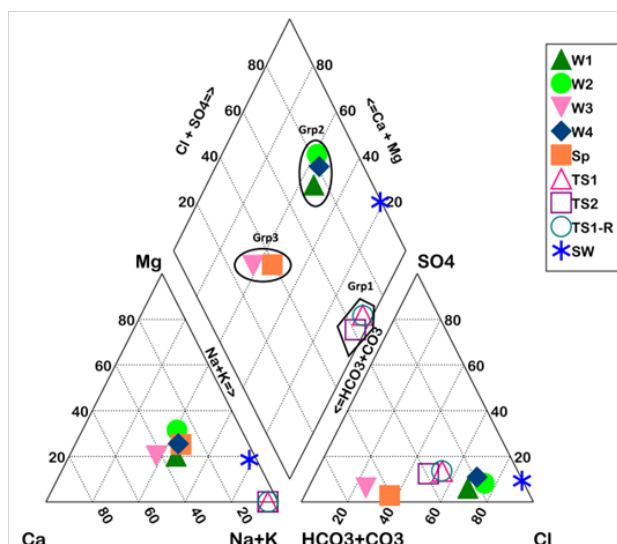


Figure 2 Piper diagram for groundwater samples.

Cation-anion concentration and mineral solubility

Major ion composition and ionic ratios can act as a track-record of water-rock interaction during flow²⁹ the concentrations of some major ions and ionic ratios were plotted against Cl, which is regarded as conservative. The Na vs Cl shows that the samples are relatively placed below or above the halite dissolution line (Figure 3A). The Na/Cl ratio (>1) (Table 3), show that the increases of Na content in Grp1 (TS₁, TS₂ and TS₁-R) reflects a progressive reaction with plagioclase feldspar. Despite its relatively slow weathering process, plagioclase feldspar (albite-anorthite) consumes acidity to produce low temperature clay mineral;³⁰ as a result, the pH in Grp1 increased and the solubility of calcite decreased; which instantly has lowered the HCO₃/Cl and Ca/Na ratios (Table 3). In addition to the precipitation of calcite, high Na concentration could be related to cation-exchange and high ionic mobility of Na with respect to Ca.³¹ The Sr vs Cl (Figure 3D) shows an increase of Sr with increasing Cl for waters in Grp2 and Grp3 ($W_2 > W_1 > W_4 > SP > W_3$, respectively). The Sr content shows a general correlation with Ca and SO₄, especially at higher concentrations, indicating association with gypsum/anhydrite.³² However, the Sr/Ca ratios ($2-7 \times 10^{-3}$) (Table 3) are in the range expected from water draining plutonic, volcanic and metamorphic rocks.³³ The molar ratio of Br/Cl may be used to characterize the source and evolution of dissolved Cl as marine and evaporite sources have characteristic signatures.^{34,35} The Br vs Cl, (Figure 3G) shows an increase of Br with increasing Cl content, and Br/Cl ratio (Table 3) reveals a relatively constant value for most samples (1.2 to 1.7×10^{-3}) suggesting a conservative behaviour

of Br/Cl during water-rock interaction. These values are close to the ocean constant molar ratio (1.5×10^{-3}), suggesting that recharged waters are reflecting the marine aerosol ratio.⁸ Consequently, recharge seems to be dominantly affected by the coastal atmospheric deposition along the western part; as towards the east the continental component becomes more prominent. Thus, in Bahia Concepcion location (Gulf of California),²⁸ shallow seawater presented higher Br/Cl molar ratio ($\approx 5 \times 10^{-3}$) due probably to the sink of Cl as result of evaporates precipitation. In addition, the importance of Br/Cl, may involve the alteration of andesitic rocks; which present a Br/Cl ratio ranging from 1.5 to 6.3×10^{-3} .³⁶ In general, samples composition is unlikely to be a result of mixing with seawater, with molar ratios of Mg/Cl, Na/Cl, Ca/Cl, SO₄/Cl and Cl/HCO₃+Cl (Table 3) being significantly different from those of seawater. Sample W₃ shows a more depleted Br/Cl ratio (0.8×10^{-3}) as compared to the other samples; suggesting dissolution or addition of Cl from an evaporate source;³⁵ or loss of Br from the fluids to diagenetic solids.³⁶ In Grp2, Mg/Ca increases with Cl increase (Figure 3B); however, the Mg/Ca ratio (Table 3) increase occurs independently of Cl in the unevolved fresh-waters of Grp3. The concentration of Li and F correlates positively with Cl increase, suggesting the presence of evolving groundwater (Figure 3F&H). The SO₄ vs Cl (Figure 3C) shows a positive correlation; however, NO₃ vs Cl (Figure 3I) shows a random distribution, suggesting that SO₄ and NO₃ patterns lack a relation to anthropogenic origin. The Ca as related to HCO₃ and SO₄ (Figure 4A&B) show a nonlinear correlation along the gypsum and calcite dissolution lines; suggesting a different source of sulphide and HCO₃ contents. The importance of Ca, and relatively high Ca/Cl ratio (Table 3) in samples W₁, W₂, S_p, W₃ & W₄ may be explained by dissolution of Na-Ca plagioclase, and calcite. The importance of Cl in samples W₁, W₂, W₄, TS₁, TS₂ & TS₁-R (HCO₃/Cl <1) may be related to Cl liberation to the fluid phase during alteration of Cl-bearing amphibole and biotite. The Ca+Mg/HCO₃ ratio (Table 3) indicates the dominance of Ca and Mg in Grp2, which could be explained by the weathering of ferromagnesian silicates, Ca-Na plagioclase and/or carbonates rocks. By contrast, waters in Grp1 and Grp3 show more HCO₃ content which may be related to the weathering of Na-rich feldspars (Na/Ca+Mg > 40 for Grp1), and/or relative recent introduction of atmospheric CO₂ into the system (Grp3) (Table 3). The concentration of NO₃ is below the World Health Organization (WHO) maximum contaminant level (MCL) (<5 mg/l) and lacks correlation with DO values (Table 1); more oxygenated water shows low values of NO₃. All chemical elements and compounds are at concentration below drinking water standards, except for total dissolved solids of samples W₂ and W₄. Dissolved oxygen shows values ranging from 1.9 to 11.3 mg/l (Table 1). These values are not representing the actual conditions at depth as water might be affected by the exposure to the atmosphere. The aqueous speciation calculation using Phreeqc (Table 4) shows that the Fe(II) species (Fe²⁺, FeHCO₃⁺, FeCO₃⁺) are dominant in most samples. However, samples TS₁ (TS₁-R duplicate), and TS₂

show the importance of insoluble $\text{Fe}(\text{OH})_3$, which explain the low concentration of dissolved Fe. For samples W_1 & W_4 the concentration of dissolved oxygen was associated with high dissolved Fe and Mn (766 and $80\mu\text{g/l}$; 5.9 and $11.5\mu\text{g/l}$, respectively); this suggests that reducing conditions may be prevailing in the groundwater supplying these wells, as supported by negative Eh field measurements (Table 1). All samples are oversaturated with respect to goethite and hematite (Table 5) which may reflect the intensity of dehydration processes at surface. The concentration of Fe^{2+} in sample W_1 seems to be controlled by siderite dissolution which provides more Fe-oxides under oxidizing conditions. Relative cation content in sampled groundwater suggests that Grp2 and Grp3 (Figure 2) are related to geological units similar in nature (granodiorite and tonalite) (Figure 1) while samples of Grp1 has travelled through contrasting geological units (basaltic, Mg-andesite and meta-sedimentary rocks).

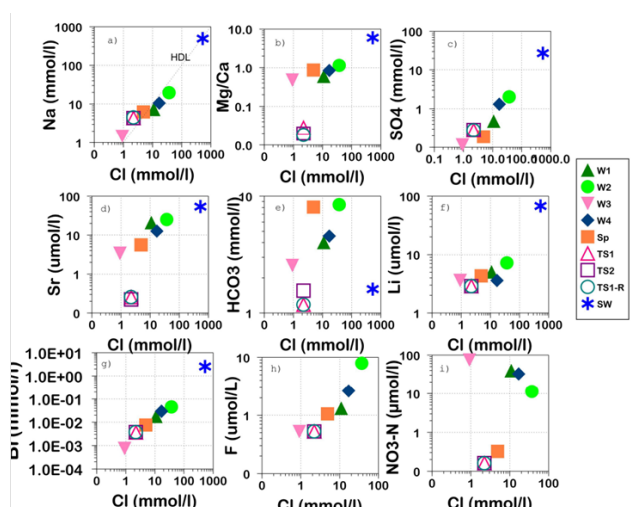


Figure 3 Logarithmic plots showing relationship of major and trace elements to Cl; HDL, Halite Dissolution Line.

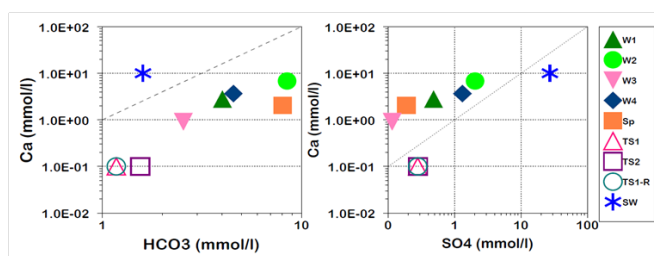


Figure 4 Logarithmic plots showing relationship between concentrations of chemical species:

- Ca vs HCO_3 , dashed line shows the Calcite Dissolution Line
- Ca vs SO_4 , dotted line shows the Congruent Gypsum Dissolution Line

Minor and trace constituents in groundwater

Minor and trace elements may be used as additional indicators of groundwater flow path and residence time. The important concentration of trace and minor constituents in Grp2 and Grp3, suggests a larger residence time as compared to thermal waters in Grp1, this argument appears to agree with the larger Li concentrations found between 35-

$50\mu\text{g/l}$ for the former samples as referred to $20\mu\text{g/l}$ for the thermal waters. The presence of other constituents suggests a conceptual support of different processes describing groundwater functioning. Total arsenic ($_{\text{As}}$) in wells W_1, W_2, W_3 & W_4 , as well as in spring S_p , with a maximum value of $7.57\mu\text{g/l}$ recorded in well W_1 is below the suggested MCL ($10\mu\text{g/l}$) for drinking water (Table 2). Despite the pH values which range between 6.78 and 7.22, the important value of $_{\text{As}}$ in samples W_1, W_2, S_p, W_3 & W_4 could be explained by the dominance of the anoxic form arsenite, $\text{As}(\text{III})$, which has more affinity for solids above pH 7-8, and which its sorption decreases with increasing ionic strength.³⁷ These results appear to confirm the origin of these samples derived from recharge-transit conditions as reported by Tóth.^{7,25} The concentration of Pb, Ni, Hg, Cd, Cu & Cr are below the proposed standard of the WHO (Table 2). Although, the concentration of Zn ($25.8\text{--}209.0\mu\text{g/l}$) are below the permissible limit (5.0 mg/l); however, this metal seems to be concentrated in samples W_1, W_2, W_3 & W_4 as compared to the rest of the samples suggesting extensive water-rock interaction. Uranium content in W_1, W_2, S_p, W_3 & W_4 shows values above the WHO recommended limit ($2\mu\text{g/l}$) (Table 2). The concentration of U in samples W_1, W_2, S_p, W_3 & W_4 with values from 4.0 to $21.5\mu\text{g/l}$ shows a relative correlation with the changing of HCO_3 content in the water. This is due to the positive association between U and pH in some environments,³⁸ as the dissolved HCO_3 increase the acidity by lowering the pH, which increases U solubility (as U^{4+} and UO_2^{2+}). By contrast, thermal waters in Grp1 (TS_1 & TS_2) show low values ($<0.005\mu\text{g/l}$) which might be related to the pH and instability of carbonates complexes at high temperatures. In the study region, the origin of U in samples W_1, W_2, S_p, W_3 & W_4 could be related to the leaching of sediments derived from granodiorite-tonalite formations under acidic conditions; groundwater remaining in contact with rocks for a long period of time have higher U contents than those with short residence.³⁷ Consequently, this residence time proxy suggests that W_2 might be supplied with the oldest groundwater. Alike uranium, V concentrations in samples W_1, W_2, S_p, W_3 & W_4 ($14.7\text{--}89.9\mu\text{g/l}$) appear to be correlated to the pH and the concentration of HCO_3 in the solution (Tables 1 & 2). The WHO has set a limit of 0.7 mg/l for Ba in drinking water. Samples W_2 and W_4 (184.0 and $187.0\mu\text{g/l}$, respectively) show the highest values of Ba, which are equivalent to high concentration of SO_4 (Tables 1 & 2). Generally, concentration of SO_4 in groundwater presents an important control on the concentration of dissolved Ba because of the low solubility of barite (BaSO_4 , $K=10^{-10}$).³⁹ The high Ba concentrations occur in groundwater that is anomalously depleted in dissolved sulphate.⁴⁰ Samples W_2 and W_4 are saturated with respect to mineral barite (Table 5), suggesting that Ba concentration is controlled by the solubility of this mineral. However, in samples W_1 and S_p , the concentration of Ba (80.5 and $127\mu\text{g/l}$, respectively) is accompanied by low levels of SO_4 which could be related to the groundwater being subjected to anoxic conditions.⁴⁰ The precipitation of barite can trigger the levels of other element which can represent health hazard at high concentration like radium (Ra). In Grp1 (TS_1, TS_2 and $\text{TS}_1\text{-R}$), high concentrations of tungsten (W) were recorded (Table 2). Many studies have shown that W can be toxic and carcinogenic.⁴¹ This high concentration in TS_1, TS_2 and $\text{TS}_1\text{-R}$ could be related to the important pH and the geothermal characteristics of groundwater. Waters in Grp1 are characterised by the presence of Gallium (Ga) and Germanium (Ge), characteristic trace elements of geothermal waters. The Ge and Ga concentration are ranging from 3.2 to $3.75\mu\text{g/l}$ and 1.2 to $1.25\mu\text{g/l}$ (Table 2), respectively.

Table 2 Trace, minor elements, and stable isotopes in groundwater samples (DL*: Detection limit)

Sample ID	Al (µg/l)	As (µg/l)	Ba (µg/l)	Br (mg/l)	Cd (µg/l)	Co (µg/l)	Cr (µg/l)	Cu (µg/l)	F (mg/l)	Fe (µg/l)	Li (µg/l)	Mn (µg/l)	Ni (µg/l)	Pb (µg/l)	Mo (µg/l)	Se (µg/l)	Sb (µg/l)	Si (µg/l)	Sr (µg/l)	Ge (µg/l)	Ga (µg/l)	Zn (µg/l)	U (µg/l)	V (µg/l)	W (µg/l)	SiO ₂ (mg/l)	H ₂ SiO ₄ (mg/l)	δ ¹⁸ O (‰)	δ ² H (‰)
DL*	2	0.03	0.1	0.03	0.01	0.005	0.5	0.2	0.01	10	1	0.1	0.3	0.01	0.1	0.2	0.01	200	0.04	0.01	0.01	0.5	0.001	0.1	0.02				
W ₁ (C)	53	7.57	80.5	1.42	0.45	0.048	3	5.5	<0.05	760	35	5.9	<2	0.75	2.9	<1	<1	26200	1830	0.55	<0.05	1190	5.04	47.7	0.5	56.05	72.856	-8.1	-58
W ₁ (C)	13	6.38	184.0	3.58	0.06	0.034	<3	3.1	<0.30	<50	50	1.0	<2	0.10	7.4	<1	<1	37200	2200	0.10	<0.05	262.0	21.50	63.6	0.1	79.58	103.44	-8.3	-58
SP Sp	21	3.85	127.0	0.60	<0.05	0.118	<3	1.1	<0.04	<50	30	4.9	<2	0.25	3.1	<1	<1	38800	489	0.05	<0.05	25.8	7.86	89.9	<0.1	83.00	107.88	-8.1	-57
W ₁ (C)	19	3.59	17.5	0.06	<0.05	<0.030	<3	<1.0	<0.02	<50	25	0.9	<2	0.10	1.1	2	2	18800	299	0.05	<0.05	123.0	4.04	50.5	0.2	40.22	52.278	-11.7	-62
W ₂ (C)	16	0.94	187.0	2.33	<0.05	0.056	<3	4.0	<0.10	80	25	11.5	<2	0.30	4.3	<1	<1	21300	1110	0.10	<0.05	209.0	4.00	14.7	<0.1	45.57	59.231	-9.0	-63
TS ₁ (TS ₁)	41	0.18	0.8	0.29	<0.05	<0.030	<3	<1.0	<0.02	<50	20	1.0	<2	0.15	4.2	<1	<1	21300	22.4	3.55	1.25	35.3	<0.005	0.5	2.45	45.57	59.231	-8.2	-57
TS ₂ (TS ₂)	35	0.16	0.9	0.30	<0.05	<0.030	<3	<1.0	<0.02	<50	20	1.9	<2	0.05	4.0	<1	<1	19700	19.5	3.20	1.30	<3.0	<0.005	0.5	2.15	42.15	54.781	-8.0	-56
TS ₁ -R (TS ₁ -R)	29	0.22	0.5	0.30	<0.05	<0.030	<3	<1.0	<0.02	<50	20	0.6	<2	0.05	4.2	<1	<1	20600	22.4	3.75	1.25	5.1	<0.005	0.5	2.35	44.07	57.284	-8.2	-57

Table 3 Molar ratios of major and trace elements

Samples	Molar ratio												
	Br/Cl	Sr/Ca	Mg/Ca	Na/Cl	Ca/Na	Mg/Cl	Ca/Cl	SO ₄ /Cl	Cl/HCO ₃ +Cl	Na/Ca+Mg	HCO ₃ /Cl	Ca+Mg/HCO ₃	ΣCO ₃ /silica
W ₁	1.62E-03	7.54	0.59	0.67	0.38	1.50E-01	2.53E-01	4.39E-02	0.73	1.66	0.37	1.101856	1.122846
W ₂	1.20E-03	3.69	1.14	0.53	0.35	2.08E-01	1.83E-01	5.46E-02	0.81	1.35	0.23	1.719453	6.965992
S _p	1.54E-03	2.72	0.86	1.28	0.33	3.63E-01	4.20E-01	3.84E-02	0.38	1.64	1.65	0.475266	2.120156
W ₃	8.19E-04	3.73	0.48	1.6	0.62	4.76E-01	9.99E-01	1.25E-01	0.26	1.08	2.78	0.529507	2.330804
W ₄	1.72E-03	3.45	0.84	0.62	0.35	1.82E-01	2.16E-01	7.73E-02	0.79	1.57	0.27	1.485217	5.526619
TS ₁	1.63E-03	2.56	0.03	1.98	0.02	1.26E-03	4.49E-02	1.25E-01	0.62	42.82	0.53	0.087233	1.296105
TS ₂	1.69E-03	2.23	0.02	1.97	0.02	8.88E-04	4.48E-02	1.27E-01	0.56	43.17	0.7	0.065623	1.129737
TS ₁ -R	1.70E-03	2.56	0.02	2.07	0.02	8.20E-04	4.52E-02	1.26E-01	0.62	44.95	0.53	0.086393	1.238819
Sea water	4.82E-03	5.51	5.95	0.92	0.02	1.11E-01	1.86E-02	5.04E-02	1	7.06	0	42.58125	

Table 4 Speciation of Fe(II), Fe(III) in sampled waters

Samples	Fe speciation								
	Fe(II)				Fe(III)				
	Fe ²⁺	FeHCO ₃	FeCO ₃	FeSO ₄	Fe ³⁺	Fe(OH) ₃	Fe(OH) ₂ ⁺	Fe(OH) ₄ ⁻	FeCl ⁺
W1	1.05E-05	2.42E-06	3.29E-07	2.27E-07	9.56E-17	2.19E-08	1.46E-08	3.12E-10	9.38E-08
W2	3.10E-07	1.12E-07	6.10E-09	1.37E-08	5.76E-18	8.14E-11	1.28E-10	5.74E-13	7.20E-09
SP	2.91E-07	1.40E-07	1.15E-08	2.58E-09	7.34E-21	2.94E-13	3.51E-13	2.32E-15	1.20E-09
W3	3.39E-07	6.34E-08	1.06E-08	3.17E-09	3.62E-17	1.96E-08	1.06E-08	3.02E-10	3.17E-10
W4	1.09E-06	2.59E-07	2.36E-08	4.87E-08	4.67E-18	2.40E-10	2.53E-10	2.29E-12	1.38E-08
TS1	9.06E-08	6.28E-09	1.58E-07	2.84E-09	5.47E-24	2.48E-08	6.85E-11	7.20E-08	2.08E-10
TS2	1.05E-10	8.85E-12	3.08E-10	2.98E-12	1.45E-23	9.97E-08	2.34E-10	3.47E-07	2.41E-13
TS1-R	9.06E-08	6.30E-09	1.58E-07	2.84E-09	5.49E-24	2.48E-08	6.84E-11	7.19E-08	2.06E-10

Table 5 Saturation Index

	Al bit e	Ca- Mon tm	Ca lci te	Chal cedo ny	Fe (OH) ₃	Gib bsit e	Geo thit e	He mati te	Illite	K- Felds par	K- Mi ca	Kao linit e	Qu art z	sid erit e	Dol omi te	Ara goni te	Ba rit e	Ha lit e	Chl orit e	Chr ysot ile	gyp su m	Wit heri te
W1	0.05	4.66	0	0.21	0.11	1.68	5.94	13.89	3.77	1.01	9.97	5.48	0.65	0.02	-0.1	-0.14	-0.18	-5.79	-2.86	-5.26	-1.91	-3.59
W2	0.09	4.12	0.19	0.34	-2.45	1.27	3.46	8.94	3.17	0.85	9	4.91	0.77	-1.69	0.58	0.05	0.41	-4.91	-3.72	-5.32	-1.24	-3.47
SP	0.26	5.12	-0.05	0.42	-4.58	1.61	1.15	4.28	4.02	1.27	10.07	5.74	0.86	-1.46	-0.09	-0.2	-0.29	-6.2	-5.08	-6.37	-2.4	-3.28
W3	-1.49	2.96	-0.47	0.06	0.05	1.2	5.89	13.78	2.08	0.11	8.12	4.22	0.5	-1.47	-1.14	-0.61	-1.19	-7.52	-5.99	-6.6	-2.74	-4.24
W4	-0.45	3.73	-0.06	0.15	-1.69	1.4	4.05	10.09	2.69	0.24	8.63	4.81	0.6	-1.14	-0.11	-0.21	0.54	-5.46	-4.64	-5.93	-1.46	-3.41
TS1	-1.65	-2.47	0.32	-0.16	-0.49	-1.04	5.72	13.5	-2.39	-1.36	2.22	-0.74	0.24	-0.24	-0.67	0.18	-2.28	-6.68	1.71	0.35	-3.32	-3.87
TS2	-1.51	-2.35	0.48	-0.16	0.5	-1	6.49	15	-2.19	-1.14	2.49	-0.63	0.26	-2.98	-0.55	0.34	-2.14	-6.67	1.01	0.08	-3.34	-3.6
TS1-R	-1.82	-2.88	0.32	-0.18	-0.49	-1.19	5.72	13.5	-2.84	-1.56	1.72	-1.07	0.22	-0.23	-0.85	0.18	-2.48	-6.67	0.41	-0.25	-3.32	-4.07

Geothermometry

Subsurface thermal reservoir temperature is a crucial parameter in evaluating groundwater travelling depth. Both, silica and cation chemical geothermometers were applied to the groundwater samples. The geothermometer of amorphous silica and β -Cristobalite (Table 6) yielded temperatures which are below the measured temperature of groundwater samples. The quartz (SiO_2 -Quartz, SiO_2 -Quartz adiabatic) solubility gives higher temperatures (94°C and 127°C) comparing to Chalcedony and α , β -Cristobalite (-1;99°C). Since silica solubility increases with temperature and pH increase, temperatures yielded by quartz adiabatic geothermometer for samples TS₁, TS₂ and TS₁-R (99,96,98°C respectively) are significant and may represent the conditions at depth (Table 6). Considering the possibility of dilution and mixing with local fresh-waters, the Na-K geothermometer can be used as well to estimate groundwater temperature in TS₁, TS₂ & TS₁-R; indeed, values yielded by Na-K⁴² (82,83,80°C respectively) and Na-K⁴³ (103,104,101°C, respectively) are close to values given by quartz adiabatic geothermometer (Table 6). Based on the chemical analyses, the fingerprint of groundwater in samples W₁, W₂, S_p, W₃ & W₄ has been notably impacted by the weathering of intermediate-mafic and metasedimentary rocks. At low temperature, mafic minerals are unstable and weather much more readily than felsic minerals. Therefore it is obvious that temperature will be less prominent than the one given for samples TS₁, TS₂ & TS₁-R. The Na-K-Ca geothermometer reflects equilibrium between Na and K-feldspars, calcite or Ca-bearing minerals and geothermal waters. This geothermometer works well for waters enriched by CO₂ and Ca. The application of Na-K-Ca geothermometer in samples W₁, S_p & W₃, has yielded values of 56, 56, and 58°C respectively, reflecting a relative

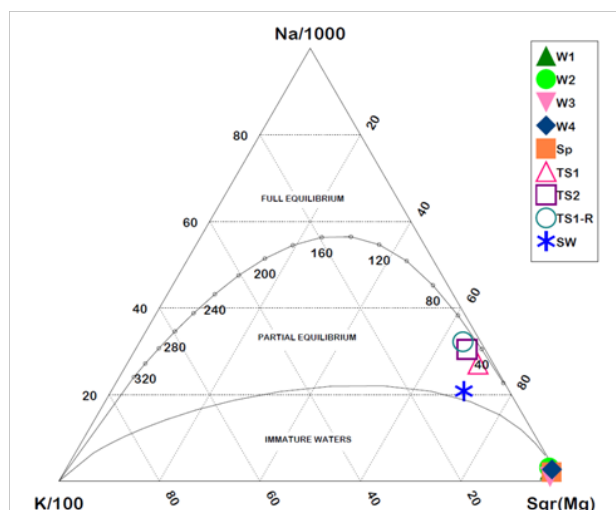
shallow travelling depth. For samples W₂ and W₄, the use of Na-K-Ca geothermometer with Mg correction is more suitable; the Mg content in these waters is important, and the use of other geothermometers can give erroneous results. This geothermometer has yielded values of 37 and 45°C for W₂ and W₄, respectively. The Na-K-Mg^{1/2} triangular diagram⁴³ provides an indication of the suitability of the waters for the application of solute geothermometers.⁴⁴ This diagram (Figure 5) shows that most samples (W₁, W₂, S_p, W₃, W₄) plot in the immature waters field indicating the interaction water-rock along the flow path and that water did not reach ionic equilibrium. The TS₁, TS₂ & TS₁-R samples present more suitability for geothermometry estimates and plot in the partially equilibrated field yielding temperatures between 80°C and 100°C. The evidence of mixing in these geothermal waters can be inferred from the relationship between silica and carbonates; when a geothermal fluid reach boiling conditions, the carbon dioxide partitions into vapour phase producing a residual liquid which is depleted in carbonate.⁴⁵ However, if dilution occurs before boiling and prevents degassing of the fluid, then the resultant discharge will retain the dissolved carbon dioxide and consequently have a high $\Sigma\text{CO}_3/\text{SiO}_2$ ratio.⁴⁵ According to the $\Sigma\text{CO}_3/\text{SiO}_2$ ratio (1.5-2.2), groundwater in TS₁, TS₂ & TS₁-R could be the result of mixing with no equilibrated water during an early stage of fluid ascent. The possibility of mixing of geothermal fluids with immature water has a negative effect on the reliability of silica geothermometers; in contrast, cation geothermometry is less sensitive to mixing and boiling processes, since it uses ratios rather than absolute abundances of the ions.⁴⁶ Therefore, temperatures given by Na-K⁴³ geothermometer and Na-K-Mg^{1/2} ternary diagram may be more informative on the sub-surface temperature in TS₁, TS₂ (and TS₁-R).

Table 6 Cation and silica geothermometers

Sample ID	SiO ₂ Fournier (1977)	A Cristobalite Fournier (1977)	β Cristobalite Fournier (1977)	Chalcedony Fournier (1977)	Quartz adiabatic Fournier (1977)	Na-K-Ca Fournier & Truesdel ⁵⁹	Na-K-Ca Mg Fournier & Potter ⁴²	Na/K Fournier & Potter ⁴²	Na/K Giggenbach ⁴³
W ₁	-9	57	10	78	107	56	56	149	168
W ₂	6	74	26	97	122	65	37	126	146
S _p	8	76	28	99	124	56	51	151	170
W ₃	-22	42	-4	61	94	58	58	270	282

Table Continued....

Sample ID	SiO ₂ Fournier (1977)	A Cristobalite Fournier (1977)	β Cristobalite Fournier (1977)	Chalcedony Fournier (1977)	Quartz adiabatic Fournier (1977)	Na-K-Ca Fournier & Truesdel ⁵⁹	Na-K-Ca Mg Fournier & Potter ⁴²	Na/K Fournier & Potter ⁴²	Na/K Giggenbach ⁴³
W ₄	-18	47	1	67	99	45	45	109	130
TS ₁	-18	47	1	67	99	64	64	82	103
TS ₂	-21	44	-2	64	96	64	64	83	104
TS _{1-R}	-19	46	-1	66	98	64	64	80	101

Figure 5 The Na-K-Mg^{1/2} diagram in mg/kg⁴³ as applied to sampled waters.

Stable isotopes: oxygen 18-deuterium

Stable and radioactive isotopes are natural tracers of water and solutes that complement geochemistry studies and provide an understanding of the origin, age, and evolution of groundwater.³⁰ Stable isotopes of water ($\delta^{18}\text{O}$, $\delta^2\text{H}$) have been used widely and successfully to distinguish between modern and palaeowaters, especially those recharged under colder conditions in the Late Pleistocene.^{9,47} In general, the isotopes dataset in Mexico are spatially and temporally deficient, with only two Global Networks for Isotopes in Precipitation (GNIP), the overall understanding of surface water-groundwater functioning over the country remains unexposed. The study of Wassenaar et al.⁴⁸ has established a small-scale groundwater isoscape for Mexico, assuming that the stable isotopes composition of shallow travelled groundwater might serve as a useful *proxy* for integrating short term (5-10 years) precipitation infiltration input. Some outcome of this study was used to analyze obtained stables isotopes data with regard to local meteoric and shallow groundwater lines. Groundwater samples in W₂, S_p, TS₁, TS₂ & TS_{1-R} (Table 2) show isotopic values that reflect recharge in lower altitudes or under warmer conditions as compared to W₃ and W₄. Consequently, samples appear to belong to different groundwater flow systems, in which the recharge conditions are likely to be variable in time and space. In general all samples plot between the Mexican Meteoric Water line $\delta D = 7.5 (\delta^{18}\text{O}) + 6.1\text{‰}$, and the shallow travelling groundwater line $\delta D = 7.9 (\delta^{18}\text{O}) + 6.4\text{‰}$ ⁴⁸ (Figure 6A). Stable isotopes values are between -11.7 and -8.0‰ for $\delta^{18}\text{O}$; and -82 to -56‰ for $\delta^2\text{H}$ (Table 2); with the most depleted

values in samples W₃ and W₄. In general, these values are depleted as compared to the tropical and summer falling rains which range between -6 to 0‰⁴⁹⁻⁵¹. The low stable isotopes in Baja California were linked to tropical cyclones which produce rain with distinctly lower $\delta^{18}\text{O}$ values (<-6‰) than other tropical rain systems.⁵² About 50% of the formed tropical cyclones are generating heavy rain (≈ 400 mm/d) in the northern part of Baja California and along the pacific coast.⁵³ The study of Jáuregui⁵⁴ has emphasized that the western Mexican coast was exposed to a greater number of hurricanes than the coast of the Gulf of México. In some studies, little evidence has been detected regarding preservation of hurricane rain as groundwater.⁵⁵ In karst areas where rapid recharge is likely to occur, the isotope signature of transient pulses of low $\delta^{18}\text{O}$ water originating as hurricane rainwater can be detected in speleothem calcite.⁵⁶ In groundwater, however, a hurricane pulse has little or no enduring effect. A possible exception may occur under unconfined conditions where values might reflect recharge of hurricane rainwater.^{55,57} Compared to groundwater sampled in 2007,⁴⁸ the isotopic signature of current samples reflect generally cooler temperatures conditions of recharge (Figure 6A). Depending on the nature of the lithological framework which is dominated by volcanic and meta-sedimentary rocks, and on the chemical patterns, which indicate a significant water-rock interaction during groundwater flow; the depleted stable isotopes values may advance new explanations regarding the relation between palaeo-climate, hurricanes effects, and groundwater isotopic signature in Northern Baja California. Further studies linking recent and palaeo-recharge processes as related to tropical cyclone events need to be conducted, as low $\delta^2\text{H}$ -values could indicate recharge during the deglaciation period, when the $\delta^2\text{H}$ of the rain water was much more negative, probably by as much as -50‰.⁵⁸ The Cl vs $\delta^{18}\text{O}$ shows that depleted values of $\delta^{18}\text{O}$ correspond to high (W₂, W₄) and low (W₃) Cl content (Figure 6B). The chemical analyses, suggest an increase in groundwater residence time in W₂ and W₄; thus, the isotopic signature might confirm cold recharge conditions dating back to the late glacial period. More depleted $\delta^{18}\text{O}$ and low Cl content in W₃ may be the result of modern recharge from tropical cyclones. In general, integrated investigations must be conducted in the study region to further understand the occurrence of groundwater with different physico-chemical and isotopic signals.

Groundwater flow systems determination

Based on the physico-chemical, isotopic evidences and geothermometry; distinct flow conditions might be identified:

Group 1: (less mineralized thermal-waters (Na-Cl-HCO₃)) is represented by TS₁, TS₂ & TS_{1-R}, these alkaline groundwaters are characterized by low TDS and high pH. Their Na/Ca+Mg, Na/Cl and HCO₃/SiO₂ ratios reflect the importance of hydrolysis of Na-rich silicates. These waters are characterized by the presence of Ga, Ge, Mo

and W. The temperature in the reservoir was estimated at 80-100°C and the HCO_3^- concentration suggests chemical modifications with cooling during water upraise. The losses of CO_2 during water ascent lower the solubility of CaCO_3 to a greater amount; which explains the saturation of the samples with respect to calcite and aragonite. Despite the geological environment of the thermal waters which is characterized by volcanic (basalt, Mg-andesite, tonalite) and meta-sedimentary rocks; the low concentration of Mg indicates the limited access to Mg-naturally occurring rocks. The depleted Sr values (19-

22.4 $\mu\text{g/l}$) could indicate a low contact with sedimentary carbonate rocks. Saturation with respect to silicates minerals (K-mica, chlorite, chrysotile and talc), and ferruginous products (goethite and hematite) could be originated from the hydrothermal weathering of amphibole-biotite forming the metamorphic rocks. The geothermometry and physico-chemical characteristics (Li,pH,T,TDS) of the thermal springs may reflect deep and rapid flow system. The nature of this flow is propose as regional.

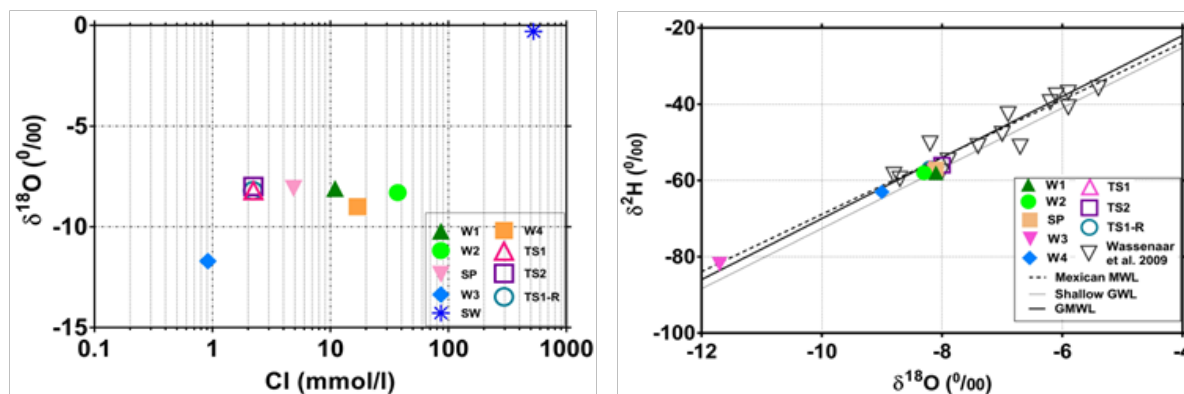


Figure 6 (A) $\delta^{18}\text{O}$ - $\delta^2\text{H}$ plot for sampled groundwater. (B) $\delta^{18}\text{O}$ vs Cl content.

Group 2: (brackish and mineralized groundwater (Na-Ca-Mg-Cl facies)) is represented by samples W_1 , W_2 and W_4 , these waters are enriched in minor and trace elements ($As, Sr, Zn, Pb \& Cu$). Groundwater temperature at depth was estimated at 37, 45 & 56°C for W_2 , W_4 & W_1 , respectively. The temperature increase was accompanied by a mineralisation decrease; this may imply that water supplying W_2 and W_4 have being in contact with mafic minerals which preferentially weather at low temperature. The presence of mafic minerals was indicated by the importance of magnesium content in W_2 and W_4 (188.0 and 75.1 mg/l, respectively). The Mg/Ca ratio (Table 3) in W_2 is greater than 0.9 suggesting that groundwater type is strongly affected by the weathering of Mg-rich silicates minerals. Whereas the Mg/Ca ratio in W_1 and W_4 show association with carbonate rocks. The bicarbonate/silica ratios (4-6 < 10) and TDS values show that groundwater in these samples is the result of silicate and carbonate weathering. The high levels of Sr and Na show evidence of an extensive water-rock interaction. Based on geothermometry, mineralisation, and chemical water age tracers (As, Sr, Li, Cl, U) these samples reflect the occurrence of long residence groundwater, and in term of flow systems this group could be classified as intermediate flow.

Group 3: (fresh and peripheral groundwater (Ca-Na-MG- HCO_3 -Cl)) is represented by S_p and W_3 , these waters are very distinctive; S_p water was sampled from a spring at 577m asl of altitude; it has a TDS of 997.49 mg/l and it presents an important HCO_3^- concentration ($\text{HCO}_3^-/\text{Cl} \approx 1.64$); the Na/Ca ratio suggests cation-exchange. Regarding minor and trace elements, S_p sample shows the importance of $As, Ba, Co, Pb, V \& U$ (Table 2), suggesting a significant water-rock interaction. The temperature at depth was estimated at 56°C indicating that waters supplying this spring were travelling in a relatively shallow depth implying a local flow system. Groundwater in W_3 represents younger characteristics; this is reflected in HCO_3^-/Cl and Ca/Na molar ratios (2.78, and 0.62, respectively) and in low value of TDS (310.23mg/l). The depth to the water-table was estimated to 78

m, thus, this water could represent recharge-transit conditions. In the absence of clear understanding of the history of groundwater isotopic signature in northern Baja California, the depletion of stable isotopes in W_3 may be related to groundwater recharge by recent hurricanes.

Conclusion

Protected areas of Northern Baja California may be subjected to natural and human stress; thus any sustainable development of the area has to consider the involvement of groundwater as an ecological agent. This work was conducted in order to determine the groundwater flow systems characterizing the protected area of Valle de Los Cirios. Results acknowledged the presence of different groundwater flow systems, three groundwater groups were differentiated. These groups have shown the presence of water with relatively long, moderate and short residence time. The groundwater geothermometry has revealed different depths and conditions of flow with Grp1 representing a deep and rapid flow through fractured igneous volcanic and metamorphic rocks; fresh groundwater flow of Grp3 appears to travel across sedimentary and volcanic sequences in relatively shallow conditions, and groundwater in Grp2 representing more evolved travelling conditions. The dominance of alkaline and shallow brackish groundwater types appears to trigger the preferential growth of desertic vegetation which might be vulnerable to a regional lowering of groundwater table. Additional groundwater studies required to be conducted in the region in order to generate a 3D groundwater flow model, with a further application of the groundwater flow systems, and links the zone of recharge with those of discharge and their related soil and vegetation.

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Conflict of interest

None.

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