

# Effective temperature and performance characteristics of heat engines

## Abstract

For a heat engine working between two heat reservoirs, a hot reservoir at high temperature  $T_H$  and a cold reservoir at low temperature  $T_L$ , an effective temperature and an effective efficiency are introduced. The effective temperature is defined as the square root of the ratio between the net work output of the heat engine  $w_{net}$  and the *SOT* function (a new function introduced in this article). The *SOT* function is defined as the negative of the cyclic integral of the heat transfer change divided by the square of the temperature. The effective efficiency of the heat engine is defined (a novel definition) as one minus the ratio between the low temperature and the effective temperature. The effective temperature and the effective efficiency are worked out in details for the Carnot heat engine and the air standard cycles (Otto, Brayton, Stirling, and Ericsson). It was found for the considered cycles that the effective temperature is given by the expression and the effective efficiency

is given by the expression  $n_{eff} = 1 - \sqrt{\frac{T_L}{T_H}}$  for all the considered cycles. The importance

of these two proposed measures is twofold: educational and they could be used as a quick tool by the designer.

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## Introduction

For a heat engine working between two heat reservoirs, a hot reservoir at high temperature  $T_H$  and a cold temperature at low temperature  $T_L$ , what is the maximal achievable efficiency by the heat engine? This question was answered by Sadi Carnot at the beginning of the 19<sup>th</sup> century by using methods of thermodynamics.<sup>1,2</sup> It was found that the maximal efficiency is limited by the Carnot efficiency

$\left( n_C = 1 - \frac{T_L}{T_H} \right)$ . It is well known that the Carnot cycle is the most

efficient cycle operating between these specified temperature limits. The Carnot cycle includes four branches: isentropic compression, isothermal heat addition, isentropic expansion and isothermal heat rejection.<sup>3</sup> The Carnot cycle is an idealized thermodynamic cycle and is not appropriate to describe real heat engines such as air standard cycles. The Otto cycle, named after Nikolaus A. Otto, is the ideal cycle for spark-ignition reciprocating engines. The Otto cycle describes the ideal behavior of spark-ignition engines, in which the piston traces four strokes (four-stroke internal combustion engine). The internal combustion engine has intake and exhaust valves. These valves are closed during compression and expansion and they are open while exchanging intake (air mixture) and exhaust (combustion gases).<sup>3</sup> The Brayton cycle, named after George Brayton, is used to describe the behavior of the reciprocating oil-burning engine that was developed by Brayton around 1870. The Brayton cycle is used today to describe the behavior of gas turbines which includes four processes: isentropic compression and expansion, and constant pressure heat addition and rejection. The working fluid is ideal gas.<sup>3</sup>

Stirling and the Ericsson engines are considered external combustion engines. That is, the energy flows to the cylinder from outside. External combustion has some advantages compared to internal combustion. Among these is the flexibility of choosing

thermal energy sources, less air pollution due to complete combustion, more efficient use of energy sources, closed cycles operation which enables choosing best working fluids such as Hydrogen and Helium. The Stirling engine includes four processes: isothermal compression, isochoric heat addition, isothermal expansion and isochoric heat rejection. The Ericsson cycle differs from the Stirling cycle by the heat addition and heat rejection processes. While in the Stirling cycle these are isochoric processes, in the Ericsson cycle they are isobaric processes.<sup>3</sup> Curzon & Ahlborn<sup>4</sup> investigated the efficiency of a heat engine at maximum power operation in finite time.<sup>5-10</sup> They considered in 1975 a model of heat engine (usually called the Curzon-Ahlborn engine) with finite heat transfer rates. The heat engine produces zero power output in the extremes of very slow operation and very fast operation. It was found that the heat engine attains a maximum power point with the efficiency at this point being one minus the square root of the ratio between the temperature of the cold reservoir and

the temperature of the hot reservoir  $\left( n_{CA} = 1 - \sqrt{\frac{T_L}{T_H}} \right)$ . Bejan<sup>11</sup>

showed in 1994 that this result of heat engine efficiency at maximum power operation was previously derived by Novikov<sup>12</sup> in 1957 when analyzing the performance of nuclear power plants.

Harvey S Leff<sup>13</sup> considered four air standard cycles (Otto, Diesel, Brayton, Atkinson) and found that the efficiency of these cycles at maximum power operation is similar to the achieved result by Curzon-Ahlborn. The methods of irreversible thermodynamics and finite thermodynamics have been used to analyze thermodynamic systems.<sup>14-32</sup> Analysis of the air standard cycles (Otto and others) could be found elsewhere.<sup>33-60</sup> In this article we ask: what are the *effective* temperature and the *effective* efficiency of a heat engine working between two heat reservoirs for the cycles described above? The answers to these questions are given in the following paragraphs. The following sections describe: the *SOT* function in section II, the

model of the heat engine in section III; the Carnot cycle is considered in section IV, the air standard cycles are considered in section V, numerical examples are given in section VI, and finally summary and conclusions are given in section VII.

### The SOT function

For a thermal process that goes from state A to state B, the negative integral of the heat transfer change divided by the square of the temperature from state A to state B is given by (a new function introduced in this article and it is called by the author the **SOT** function):

$$SOT = -\int_A^B \frac{\delta q}{T^2} \tag{1}$$

For isothermal process the **SOT** function is given by:

$$SOT = -\frac{1}{T^2} \int_A^B \delta q \tag{2}$$

The first law of thermodynamics for a process with ideal gas working fluid is given by:

$$\delta q = dE + pdV \tag{3}$$

By considering the various process types and using the first law of thermodynamics (eq. 3), makes it easy to write explicit forms of the **SOT** function. In the following subsections some typical processes are considered.

#### Constant volume process

For the constant volume process ( $v=\text{const.}$ ), the heat transfer change is proportional to the temperature change ( $dT$ ) with constant volume heat capacity proportionality factor ( $c_v$ ). In this case the **SOT** function is given by:

$$SOT = c_v \left( \frac{1}{T_A} - \frac{1}{T_B} \right) \tag{4}$$

#### Constant pressure process

For the constant pressure process ( $p=\text{const.}$ ), the heat transfer change is proportional to the temperature change ( $dT$ ) with constant pressure heat capacity proportionality factor ( $c_p$ ). In this case the **SOT** function is given by:

$$SOT = c_p \left( \frac{1}{T_A} - \frac{1}{T_B} \right) \tag{5}$$

#### Constant temperature process

For the constant temperature process ( $T=\text{const.}$ ), the heat transfer change is proportional to the volume change ( $dv$ ) with pressure proportionality factor ( $p$ ). In this case the **SOT** function is given by:

$$SOT = \frac{R}{T} \ln \left( \frac{v_A}{v_B} \right) \tag{6}$$

#### Isentropic process

For the isentropic process-constant entropy process ( $s=\text{const.}$ ), the heat transfer change is zero. In this case the **SOT** function is given by:

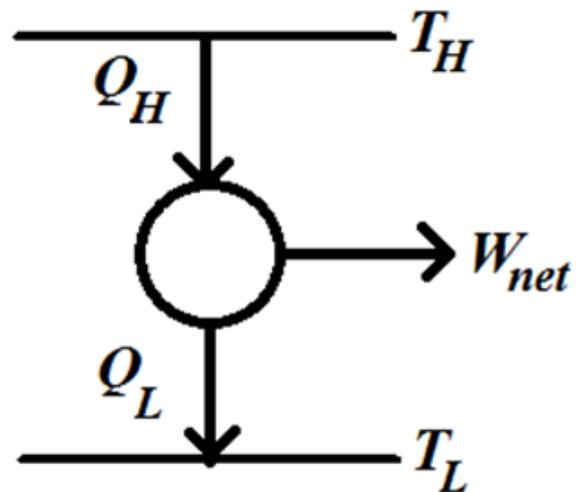
$$SOT = 0 \tag{7}$$

For a heat engine going in a cycle the **SOT** function is given by:

$$SOT = -\oint \frac{\delta q}{T^2} \tag{8}$$

### Effective temperature and efficiency of heat engines

Consider a heat engine working between two heat reservoirs, a hot reservoir at high temperature ( $T_H$ ) and a cold reservoir at low temperature ( $T_L$ ). The schematics of the heat engine are depicted in Figure 1.



**Figure 1** Schematics of the engine working between two heat reservoirs, a hot reservoir at high temperature  $T_H$  and a cold reservoir at low temperature  $T_L$ . The heat input to the heat engine is  $Q_H$  and the heat rejection from the heat engine is  $Q_L$ . The net work output  $w_{net}$  is given by the  $Q_H - Q_L$ .

The net work output  $w_{net}$  is calculated by means of the first law of thermodynamics (eq. 3) and is given by:

$$w_{net} = Q_H - Q_L \tag{9}$$

The **SOT** function is given by equation (8).

The effective temperature is defined as the square root of the ratio between the net work output and the **SOT** function, and is given by:

$$T_{eff} = \sqrt{\frac{w_{net}}{SOT}} \tag{10}$$

Finally, the effective efficiency is defined as the difference between unity minus the ratio between the low temperature and the effective temperature, and is given by:

$$\eta_{eff} = 1 - \frac{T_L}{T_{eff}} \tag{11}$$

#### The Carnot heat engine

The schematic of the Carnot heat engine is depicted in Figure 1. The net work output is given by (eq. 9). The Carnot cycle includes two processes: heat addition and heat rejection at constant temperatures

(high and low) connected with two isentropic (reversible adiabatic) processes. The entropy of the heat engine is calculated by means of the second law of thermodynamics and is given by:

$$\oint \frac{\delta q}{T} = 0 \tag{12}$$

By applying the assumptions of the Carnot cycle, equation (12) could be written explicitly as follows:

$$\frac{Q_H}{T_H} = \frac{Q_L}{T_L} \tag{13}$$

The SOT function for the Carnot cycle is derived by means of equation 1 and is given by:

$$SOT = \frac{Q_L}{T_L^2} - \frac{Q_H}{T_H^2} \tag{14}$$

Applying equation (13) to equation (4) aids to simplify the expression for the **SOT** function. After algebraic manipulation and using equation (9), the **SOT** function of the Carnot heat engine simplifies to the following:

$$SOT = \frac{w_{net}}{T_L T_H} \tag{15}$$

Then, the effective temperature is defined as the geometric mean of the high temperature and the low temperature, a result suggested by (eq. 15) and is given by:

$$n_{eff} = \sqrt{T_L T_H} = \sqrt{\frac{w_{net}}{SOT}} \tag{16}$$

Finally, the effective efficiency for the Carnot heat engine is given by:

$$n_{eff} = 1 - \frac{T_L}{T_{eff}} = 1 - \sqrt{\frac{T_L}{T_H}} \tag{17}$$

## Air standard cycles

### Ideal Otto cycle

The ideal Otto cycle is used to estimate the efficiency of spark ignition (SI) engine (Otto engine). The schematic of the Pressure–Volume (P-V) diagram of the ideal Otto cycle is shown in Figure 2. The cycle includes four processes: 1→2 isentropic compression, 2→3 constant volume heat addition, 3→4 isentropic expansion, and 4→1 constant volume heat rejection.

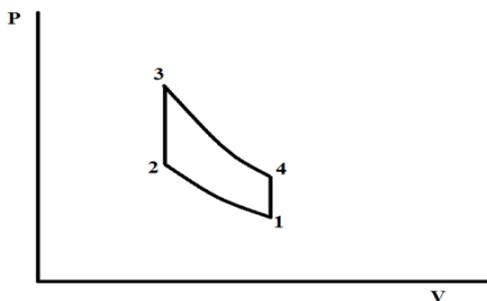


Figure 2 Schematics if the ideal Otto cycle.

For a given initial state (pressure  $P_1$ , volume  $V_1$ , temperature  $T_1$ ) with known highest temperature  $T_3$ , the application of the thermodynamic state relations and the first law of thermodynamics of the different branches lead to the following results:

The heat input to the engine  $Q_H$  is given by:

$$Q_H = c_v (T_3 - T_2) = c_v T_3 (1 - \tau a) \tag{18}$$

Where  $c_v$  is the constant volume heat capacity,  $\tau$  is the ratio between  $T_1$  and  $T_3$  and  $a$  ( $a = r^{k-1}$ ) is the compression ratio  $r$  raised to the power  $(k-1)$  with  $k$  equals the ratio between constant pressure heat capacity ( $c_p$ ) and constant volume heat capacity ( $k=c_p/c_v$ ).<sup>1</sup>

The heat rejection from the engine is given by:

$$Q_L = c_v (T_4 - T_1) = c_v T_3 \frac{(1-\tau a)}{a} \tag{19}$$

Where  $T_4$  is the temperature at state 4.

The net work output ( $w_{net}$ ) extracted by the engine is given by:

$$w_{net} = Q_H - Q_L = c_v T_3 (1 - \tau a) \left( 1 - \frac{1}{a} \right) \tag{20}$$

The **SOT** function is calculated based on equation (8). Noting that the contributions of the isentropic branches to the **SOT** function are zero, the resulting **SOT** function for the ideal OTTO cycle is given by:

$$SOT_{Otto} = \frac{c_v}{T_1} (1 - \tau a) \left( 1 - \frac{1}{a} \right) \tag{21}$$

By comparing equations (20) and (21) it is observed that the effective temperature could be deduced and is given by:

$$T_{eff} = \sqrt{\frac{w_{net}}{SOT_{Otto}}} = \sqrt{T_1 T_3} \tag{22}$$

The effective efficiency for the ideal OTTO cycle is given by equation (17), where  $T_1$  is replaced by  $T_L$  and  $T_3$  is replaced by  $T_H$ .

### Ideal Brayton cycle

The ideal Brayton cycle is used to estimate the efficiency of gas turbines. The schematic of the Pressure–Volume (P-V) diagram of the ideal Brayton cycle is shown in Figure 3. The Brayton cycle includes four processes: 1→2 isentropic compression, 2→3 constant pressure heat addition, 3→4 isentropic expansion, and 4→1 constant pressure heat rejection.

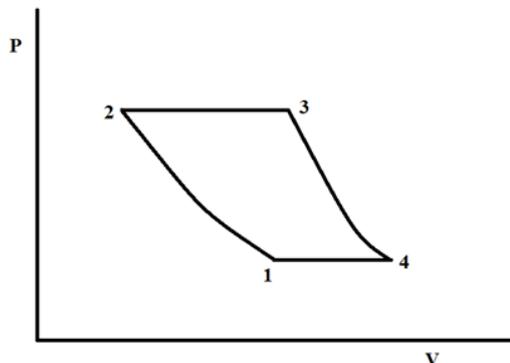


Figure 3 Schematics if the ideal Brayton cycle.

For a given initial state (pressure  $P_1$ , volume  $V_1$ , temperature  $T_1$ ) with known highest temperature  $T_3$ , the application of the thermodynamic state relations and the first law of thermodynamics of the different branches lead to the following results:

The heat input to the engine  $Q_H$  is given by:

$$Q_H = c_p(T_3 - T_2) = c_p T_3(1 - \tau a) \quad (23)$$

Where  $c_p$  is the constant pressure heat capacity,  $\tau$  is the ratio between  $T_1$  and  $T_3$  and  $a$  ( $a = r^{k-1}$ ) is the compression ratio  $r$  raised to the power  $(k-1)$  with  $k$  equals the ratio between constant pressure heat capacity ( $c_p$ ) and constant volume heat capacity ( $k=c_p/c_v$ ).

The heat rejection from the Brayton cycle is given by:

$$Q_L = c_p(T_4 - T_1) = c_p T_3 \frac{(1-\tau a)}{a} \quad (24)$$

Where  $T_4$  is the temperature at state 4.

The net work output ( $w_{net}$ ) extracted by the Brayton cycle is given by:

$$w_{net} = Q_H - Q_L = c_p T_3(1 - \tau a) \left(1 - \frac{1}{a}\right) \quad (25)$$

The SOT function is calculated based on equation (8). Noting that the contributions of the isentropic branches to the SOT function are zero, the resulting SOT function for the ideal Brayton cycle is given by:

$$SOT_{Brayton} = \frac{c_p}{T_1} (1 - \tau a) \left(1 - \frac{1}{a}\right) \quad (26)$$

By comparing equations (25) and (26) it is observed that the effective temperature could be deduced and is given by:

$$T_{eff} = \sqrt{\frac{w_{net}}{SOT_{Brayton}}} = \sqrt{T_1 T_3} \quad (27)$$

The effective efficiency the ideal Brayton cycle is given by equation (17), where  $T_1$  is replaced by  $T_L$  and  $T_3$  is replaced by  $T_H$ .

### Ideal Stirling cycle

The ideal Stirling cycle is used to estimate the efficiency of Stirling engine. The schematic of the Pressure–Volume (P-V) diagram of the ideal Stirling cycle is shown in Figure 4. The cycle includes four processes: 1→2 isothermal compression, 2→3 constant volume heat addition, 3→4 isothermal expansion, and 4→1 constant volume heat rejection.

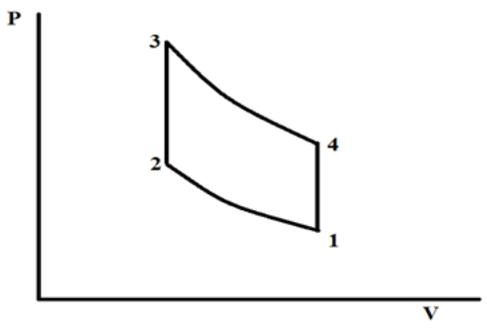


Figure 4 Schematics if the ideal Stirling cycle.

For a given initial state (pressure  $P_1$ , volume  $V_1$ , temperature  $T_1$ ) with known highest temperature  $T_3$ , the application of the thermodynamic state relations and the first law of thermodynamics of the different branches lead to the following results:

The heat input to the Stirling cycle  $Q_H$  is given by:

$$Q_H = c_v(T_3 - T_2) + RT_3 \ln(r) \quad (28)$$

Where  $c_v$  is the constant volume heat capacity,  $\tau$  is the ratio between  $T_1$  and  $T_3$  and  $a$  ( $a = r^{k-1}$ ) is the compression ratio  $r$  raised to the power  $(k-1)$  with  $k$  equals the ratio between constant pressure heat capacity ( $c_p$ ) and constant volume heat capacity ( $k=c_p/c_v$ ).

The heat rejection from the Stirling cycle is given by:

$$Q_L = c_v(T_4 - T_1) + RT_1 \ln(r) \quad (29)$$

Where  $T_4$  is the temperature at state 4 and  $R$  is the ideal gas constant.

The net work output ( $w_{net}$ ) extracted by the Stirling cycle is given by:

$$w_{net} = Q_H - Q_L = c_v RT_3(1 - \tau) \ln(r) \quad (30)$$

The **SOT** function is calculated based on equation (8). Noting that the contributions of the isothermal branches to the **SOT** function cancel each other, the resulting **SOT** function for the ideal Stirling cycle is given by:

$$SOT_{Stirling} = R \ln(r) \left(\frac{1}{T_1} - \frac{1}{T_3}\right) = \frac{R \ln(r)}{T_1} (1 - \tau) \quad (31)$$

By comparing equations (30) and (31) it is observed that the effective temperature could be deduced and is given by:

$$T_{eff} = \sqrt{\frac{w_{net}}{SOT_{Stirling}}} = \sqrt{T_1 T_3} \quad (32)$$

The effective efficiency the ideal OTTO cycle is given by equation (17), where  $T_1$  is replaced by  $T_L$  and  $T_3$  is replaced by  $T_H$ .

### Ideal Ericsson cycle

The ideal Ericsson cycle is used to estimate the efficiency of the Ericsson engine. The schematics of the Pressure–Volume (P-V) diagram of the ideal Ericsson cycle is shown in Figure 5. The cycle includes four processes: 1→2 isothermal compression, 2→3 constant pressure heat addition, 3→4 isothermal expansion, and 4→1 constant pressure heat rejection.

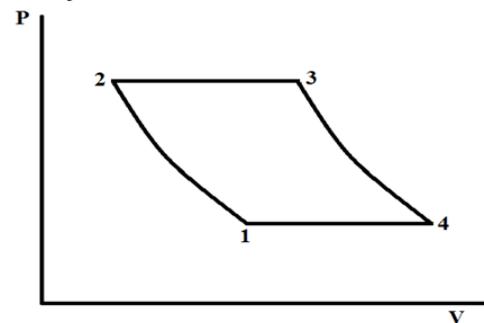


Figure 5 Schematics if the ideal Ericsson cycle.

For a given initial state (pressure  $P_p$ , volume  $V_p$ , temperature  $T_p$ ) with known highest temperature  $T_3$ , the application of the thermodynamic state relations and the first law of thermodynamics of the different branches lead to the following results:

The heat input to the engine  $Q_H$  is given by:

$$Q_H = c_p(T_3 - T_2) + RT_3 \ln(r) \tag{33}$$

Where  $c_v$  is the constant volume heat capacity,  $\tau$  is the ratio between  $T_1$  and  $T_3$  and  $a$  ( $a = r^{k-1}$ ) is the compression ratio  $r$  raised to the power  $(k-1)$  with  $k$  equals the ratio between constant pressure heat capacity ( $c_p$ ) and constant volume heat capacity ( $k=c_p/c_v$ ).

The heat rejection from the engine is given by:

$$Q_L = c_p(T_4 - T_1) + RT_1 \ln(r) \tag{34}$$

Where  $T_4$  is the temperature at state 4.

The net work output ( $w_{net}$ ) extracted by the engine is given by:

$$w_{net} = Q_H - Q_L = RT_3(1 - \tau) \ln(r) \tag{35}$$

The **SOT** function is calculated based on equation (8). Noting that

the contributions of the isobaric branches to the **SOT** function cancel each other, the resulting **SOT** function for the ideal Ericsson cycle is given by:

$$SOT_{Ericsson} = R \ln(r) \left( \frac{1}{T_1} - \frac{1}{T_3} \right) = \frac{R \ln(r)}{T_1} (1 - \tau) \tag{36}$$

By comparing equations (35) and (36) it is observed that the effective temperature could be deduced and is given by:

$$T_{eff} = \sqrt{\frac{w_{net}}{SOT_{Ericsson}}} = \sqrt{T_1 T_3} \tag{37}$$

The effective efficiency the ideal OTTO cycle is given by equation (17). where  $T_1$  is replaced by  $T_L$  and  $T_3$  is replaced by  $T_H$ .

### Numerical results

The following table compares the effective efficiency with Carnot efficiency and with the observed efficiency for some real plants. As can be seen, the effective efficiency is much closer to the observed data.<sup>60</sup>

Power Source	TL °C	TH °C	Carnot efficiency	Effective efficiency	Observed efficiency	Effective temperature K
West Thurrock (UK) coal fired power plant	25	565	64.40%	40.40%	36%	500
CANDU (Canada) nuclear power plant	25	300	48%	27.90%	30%	413
Larderello (Italy) geothermal power plant	80	250	32.50%	17.80%	16%	430
Stirling/Ericsson	27	600		41.4		512
Otto	27	1800	85.5	62		789
Brayton	27	1100	78.1	53.2		642

### Summary and conclusion

For a heat engine working between two heat reservoirs, a hot reservoir at high temperature  $T_H$  and a cold reservoir at a low temperature  $T_L$  the Carnot cycle and Curzon-Ahlborn heat engine were shortly reviewed and their performance efficiencies are given via the Carnot efficiency and the Curzon-Ahlborn respectively. The new terms **SOT** function, effective temperature and the effective efficiency were introduced. The **SOT** function was defined as minus the cyclic integral of the heat change divided by the temperature squared. The effective temperature was defined as the square root of the ratio between the net work output and the **SOT** function, and finally the effective efficiency was defined as one minus the ratio between the cold reservoir temperature and the effective temperature. The **SOT** function was calculated for different thermodynamic processes: isochoric, isobaric, isothermal and isentropic. A model of heat engine was considered and general expressions of the **SOT** function, effective temperature and effective efficiency were given as useful tools for the designer. These expressions were applied to different thermodynamic cycles: The Carnot cycle and air standard cycles (Otto, Brayton, Stirling and Ericsson). The effective temperature for the considered heat engines was given as the square root of the product of the reservoirs' temperatures and the efficiency was given as one minus the square root of the ratio between the cold reservoir temperature and the hot reservoir temperature. The derived expressions can serve two important purposes: in education and for a quick estimation tool for heat engine designers.

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### Conflict of interest

The author declares that there is no conflict of interest regarding the publication of this paper.

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